

### Types of Chemical Reactions Worksheet #1

#### **11. Chemistry**

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1)  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$  **Synthesis**
  - 2)  $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$  **Single Replacement**
  - 3)  $\text{HNO}_3 + \text{AlCl}_3 \rightarrow \text{Al}(\text{OH})_3 + 3\text{HCl}$  **Double Replacement**
  - 4)  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$  **Synthesis**
  - \*5)  $\text{Ca}(\text{OH})_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$  **Double Replacement**
  - 6)  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$  **Decomposition**
  - 7)  $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$  **Single Replacement**
  - 8)  $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$  **Double Replacement**
  - 9)  $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$  **Synthesis**
  - \*10)  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$  **Single Replacement**
- \*Remember: "H<sub>2</sub>O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 1.1) Double replacement:  $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 1.2) Single replacement:  $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 1.3) Decomposition:  $\text{Mg}(\text{ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 1.4) Synthesis:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 1.5) Double replacement:  $3\text{Cu}(\text{OH})_2 + 2\text{FeCl}_3 \rightarrow 2\text{Fe}(\text{OH})_3 + 3\text{CuCl}_2$
- 1.6) Single replacement:  $\text{Fe} + \text{Cu}(\text{NO}_3)_2 \rightarrow \text{Fe}(\text{NO}_3)_2 + \text{Cu}$  [Assume Fe<sup>2+</sup>]
- 1.7) Decomposition:  $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 1.8) Synthesis:  $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 1.9) Double replacement:  $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 2.0) Single replacement:  $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$  [Copper (II) is used here]

# The Nature Of Chemical Reactions Answers

**Lei Shi**



## **The Nature Of Chemical Reactions Answers:**

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