

Concentration Practice Problems

% Solution

1. What is the concentration (% solution) of a solution that contains 0.78 grams of MgCl_2 dissolved in 10. grams of water?

$$\% = \frac{\text{mass of solute}}{\text{mass of sol'n}} = \frac{0.78 \text{ g}}{10.78 \text{ g}} = 7.24\%$$

2. How many grams of solid are needed to prepare a 10.0% NaCl solution in 200. grams of water?

$$0.10 = \frac{x}{x+200}$$

$$0.10(x+200) = x$$

$$0.10x + 20 = x$$

$$-0.10x \quad -0.10x$$

$$20 = 0.9x$$

$$\frac{20}{0.9} = \frac{0.9x}{0.9}$$

$$x = 22.2 \text{ g NaCl}$$

3. What is the % solution of a mixture that contains 0.0500 moles of NaI dissolved in 25.0 g of water?

$$0.0500 \text{ moles NaI} \times \frac{150.0 \text{ g}}{1 \text{ mole}} = 7.5 \text{ g NaI}$$

$$\frac{7.5 \text{ g NaI}}{(7.5 + 25) \text{ g Sol'n}} = 23.1\%$$

Molarity

4. What is the molarity of a solution that contains 85.0 g of iodine (I_2) dissolved in 250. mL of ethanol solution?

$$\frac{85.0 \text{ g}}{1} \times \frac{1 \text{ mole}}{254 \text{ g}} = 0.335 \text{ mole}$$

$$M = \frac{0.335 \text{ moles}}{0.250 \text{ L}} = 1.34 \text{ M}$$

5. How many grams of $\text{Ca}(\text{NO}_3)_2$ are need to produce 1500. mL of a 0.75 M solution?

$$M = \frac{\text{mol}}{\text{L}}$$

$$0.75 = \frac{x}{1.500 \text{ L}}$$

$$x = 1.125 \text{ mol}$$

$$\frac{1.125 \text{ mol}}{1} \times \frac{164.1 \text{ g}}{1 \text{ mol}} = 184.6 \text{ g Ca}(\text{NO}_3)_2$$

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Steven M. Rice



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