Mole Conversions Practice Answer Key

Molar Mass - Calculate the Molar Mass for the following.

What is the molar mass of Magnesium metal?

24.305 gmal

Calculate the molar mass of NaCl

Nq:20.990 x1 = 20.990 C1:35.453 x1 = 35.45 3

3. What is the molar mass of NaOH?

Na: 1 x 22.990 = 22.990 D: 1 x 15.999 = 15.999 H: 1 x 1.008 = 1.008

Conversions - Calculate the following the conversions. Use your Mole Map to help you!

Determine the number of atoms in 0.50 moles He.

= 13.01 x 10 23 atoms He

Determine the number of moles in 80.00 g O₂

0 2 x 15 999 = 31 998 201

=> 1 mol = 31.9980

Determine the volume of 0.24 mol of NO₂.

4. Determine the number of moles in 7.5 x 1019 formula units of KNOs

75×1019 f.u. KNO, 6.03 ×1023 f.u KNO

Determine the number of moles in 202,22a KNO₃

K 1x 39.098 = 39.098 0: 3x 15.999 = 47.997) 101.102 g/mol N 17 14.007 = 14.007

6. Determine the number of moles in 32.5L of N₂ at STP In a = 101 102 g

Determine the number of moles in 3.4 x 10²⁶ formula units of NaCl.

34 x 10 F. U. Nacl -6.02 ×1023 f.4 Nacl

Determine the moles in 36.5L of carbon monoxide

36.5 L CO x | Imel CO

The Mole Chapter Answer Key

Linda A. LaCharity, Candice K. Kumagai, Barbara Bartz

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conventional th modynamic methods are not sufficient for the accurate treatment of these mixtures Last but not least these systems are of interest for the pharmaceutical biomedical and related ind tries In order to meet the thermodynamic challenges involved in these complex mixtures we employed a variety of traditional methods but also new methods such as the fluctuation t ory of Kirkwood and Buff and ab initio quantum mechanical techniques The Kirkwood Buff KB theory is a rigorous formalism which is free of any of the proximations usually used in the thermodynamic treatment of multicomponent systems This theory appears to be very fruitful when applied to the above mentioned difficult systems

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