Name: Key

The number of moles of substances, when compared and reduced to the lowest terms. yields a mole ratio, as shown in the sample below. Indicate the mole ratio for each of the following equations:

$$4H_2 + 2O_2 \Rightarrow 4H_2O$$
 2: 1: 2

  $H_2 + S \Rightarrow H_2S$ 
 1: 1: 1

  $4AI + 3O_2 \Rightarrow 2 AI_2O_3$ 
 4: 3: 2

  $3V_1 + 6I_1CI \Rightarrow 3V_1CI_2 + 3H_2$ 
 1: 2: 1: 1

  $SO_2 + \frac{1}{2}O_2 \Rightarrow SO_3$ 
 2: 1: 2

  $6H_2 + 3C \Rightarrow 3CH_4$ 
 2: 1: 1

FW= 17.00/mol

Convert the following masses to moles by using the information on your periodic table.

a) 
$$34.0 \text{ g NH}_3$$
 FW = 17.05 feet  
 $34.0 \text{ g (} \frac{\text{I mol}}{17.0 \text{ g}} \text{)} = 2.00 \text{ mol}$ 

b)  $51.4 \text{ g Na}_2\text{O}$  FW = 62.09 fmol

 $51.4 \text{ g (} \frac{\text{I mol}}{62.0 \text{ g}} \text{)} = 0.829 \text{ mol}$ 

c)  $4.50 \text{ g II}_2\text{SO}_4$  FW = 98.1 g fmol

 $4.50 \text{ g (} \frac{\text{I mol}}{98.1 \text{ g}} \text{)} = 0.0459 \text{ mol}$ 

d)  $229 \text{ g Ca(NO}_3)_2$  FW = 164.19 fmol

 $229 \text{ g (} \frac{\text{I mol}}{164.19} \text{)} = 1.40 \text{ mol}$ 

c)  $79.5 \text{ g NH}_4\text{OH}$  FW =  $35.0 \text{ g fmol}$ 
 $79.5 \text{ g (} \frac{\text{I mol}}{35.0 \text{ g (}} \text{)} = 2.27 \text{ mol}$ 

Convert the following to grams.

34.0 g NH1

a) 
$$17.1 \text{ moles } H_2O = 18.09 \text{ mol}$$
  $308 \text{ g}$ .

 $17.1 \text{ mol} \times \left(\frac{18.09}{1 \text{ mol}}\right) = 3089$ 

# **Skill Practice 1 Number Practice Answers Chemistry**

**Mary Lee** 

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#### **Table of Contents Skill Practice 1 Number Practice Answers Chemistry**

- 1. Understanding the eBook Skill Practice 1 Number Practice Answers Chemistry
  - The Rise of Digital Reading Skill Practice 1 Number Practice Answers Chemistry
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Skill Practice 1 Number Practice Answers Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Skill Practice 1 Number Practice Answers Chemistry
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Skill Practice 1 Number Practice Answers Chemistry

- Personalized Recommendations
- Skill Practice 1 Number Practice Answers Chemistry User Reviews and Ratings
- Skill Practice 1 Number Practice Answers Chemistry and Bestseller Lists
- 5. Accessing Skill Practice 1 Number Practice Answers Chemistry Free and Paid eBooks
  - Skill Practice 1 Number Practice Answers Chemistry Public Domain eBooks
  - Skill Practice 1 Number Practice Answers Chemistry eBook Subscription Services
  - Skill Practice 1 Number Practice Answers Chemistry Budget-Friendly Options
- 6. Navigating Skill Practice 1 Number Practice Answers Chemistry eBook Formats
  - o ePub, PDF, MOBI, and More
  - Skill Practice 1 Number Practice Answers Chemistry Compatibility with Devices
  - Skill Practice 1 Number Practice Answers Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Skill Practice 1 Number Practice Answers Chemistry
  - Highlighting and Note-Taking Skill Practice 1 Number Practice Answers Chemistry
  - Interactive Elements Skill Practice 1 Number Practice Answers Chemistry
- 8. Staying Engaged with Skill Practice 1 Number Practice Answers Chemistry
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Skill Practice 1 Number Practice Answers Chemistry
- 9. Balancing eBooks and Physical Books Skill Practice 1 Number Practice Answers Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Skill Practice 1 Number Practice Answers Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Skill Practice 1 Number Practice Answers Chemistry
  - Setting Reading Goals Skill Practice 1 Number Practice Answers Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Skill Practice 1 Number Practice Answers Chemistry

- Fact-Checking eBook Content of Skill Practice 1 Number Practice Answers Chemistry
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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