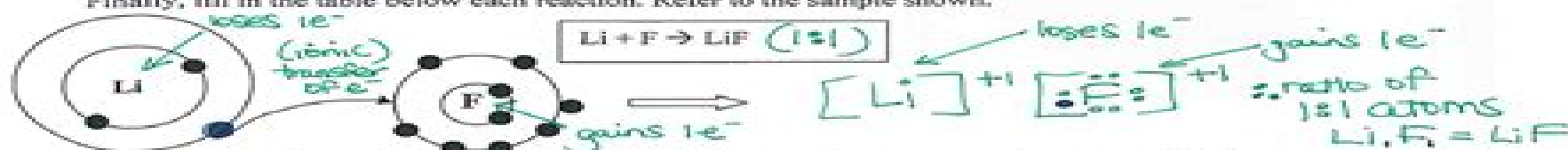


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### Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



| Atoms | Valence electrons | Electron transfer from/to each atom | Ions formed in the product |
|-------|-------------------|-------------------------------------|----------------------------|
| Li    | 1 ve              | transfers $1e^-$ to F               | $Li^+$                     |
| F     | 7 ve              | accepts $1e^-$ from Li              | $F^-$                      |

| Reactions                                                                                          | Atoms | Valence electrons | Electron transfer from/to each atom | Ions formed in the product |
|----------------------------------------------------------------------------------------------------|-------|-------------------|-------------------------------------|----------------------------|
| 1) $Li + Cl \Rightarrow LiCl$<br>$Li + \cdot\ddot{Cl}\cdot \rightarrow [Li]^+ [\ddot{Cl}]^-$       | Li    | 1 ve              | donates $1e^-$                      | $Li^+$                     |
|                                                                                                    | Cl    | 1 ve              | accepts $1e^-$                      | $Cl^-$                     |
| 2) $Ca + O \Rightarrow CaO$<br>$Ca + \cdot\ddot{O}\cdot \rightarrow [Ca]^{+2} [\ddot{O}]^{2-}$     | Ca    | 2 ve              | donates $2e^-$                      | $Ca^{+2}$                  |
|                                                                                                    | O     | 6 ve              | accepts $2e^-$                      | $O^{2-}$                   |
| 3) $Be + F \Rightarrow BeF_2$<br>$Be + 2\cdot\ddot{F}\cdot \rightarrow [Be]^{+2} [2\ddot{F}]^{2-}$ | Be    | 2 ve              | donates $2e^-$ to each F            | $Be^{+2}$                  |
|                                                                                                    | F     | 7 ve              | each F accepts $1e^-$ each          | $F^-$                      |
| 4) $Mg + S \Rightarrow MgS$<br>$Mg + \cdot\ddot{S}\cdot \rightarrow [Mg]^{+2} [\ddot{S}]^{2-}$     | Mg    | 2 ve              | donates $2e^-$ to S                 | $Mg^{+2}$                  |
|                                                                                                    | S     | 6 ve              | accepts $2e^-$                      | $S^{2-}$                   |
| 5) $K + F \Rightarrow KF$<br>$K + \cdot\ddot{F}\cdot \rightarrow [K]^+ [\ddot{F}]^-$               | K     | 1 ve              | donates $1e^-$                      | $K^+$                      |
|                                                                                                    | F     | 7 ve              | accepts $1e^-$                      | $F^-$                      |

# Skill Practice 19 Ionic Bonding Practice Answer

**Ralph Thomas Myers**



## Skill Practice 19 Ionic Bonding Practice Answer:

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African continent and with specialists from other continents It considers all areas and levels of chemistry education gives specific attention to known major challenges and encourages explorations of novel approaches The chapters in this book describe new teaching approaches approach explorations and in class activities analyse educational challenges and possible ways of addressing them and explore cross discipline possibilities and their potential benefits for chemistry education This makes the volume an up to date compendium for chemistry educators and educational researchers worldwide Glencoe Science McGraw-Hill Staff,2001-06 *TEXES Core Subjects 4-8 (211) Book + Online* Ann M.L. Cavallo,Mary D. Curtis,Peggy Semingson,Karen Allmond,Melissa Hulings,Christina Gawlik,2019-04-30 Teacher candidates seeking certification to teach the middle level grades in Texas s public schools must pass the TExES Core Subjects 4 8 exam Written by a team of faculty experts led by Dr Ann M L Cavallo Associate Dean for Research and Graduate Studies at the University of Texas at Arlington REAs test prep provides extensive coverage of the four core subject areas tested on the exam English Language Arts and Reading 806 Mathematics 807 Social Studies 808 and Science 809 In addition to a thorough review this test prep features a diagnostic test and 2 full length practice test batteries 1 in the book and 1 online at the REA Study Center that deal with every question type competency and skill tested on the exam REAs online tests run under timed conditions and provide automatic scoring and diagnostic feedback on every question to help teacher candidates zero in on the topics that give them trouble now so they can succeed on test day Amazon com **Chemistry** Katherine J. Denniston,Joseph J. Topping,Robert L. Caret,2006-05 Portland Transcript ,1860 **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science** ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction *Prentice Hall Chemistry* ,2000 **The Building News and Engineering Journal** ,1900 **School Library Journal** ,1987 **"The" Illustrated London News** ,1850 **The Case for Basic Education** James D. Koerner,1959

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