

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative
positive
electron

noble gas
atoms
ion

cation
anion
electron

electron configuration
polyatomic
empirical formula

1. An ionic bond is an attraction between oppositely charged _____. **ions**
2. Anionic bond is _____. **negative** change.
3. An ionic bond is formed by losing or gaining _____. **electrons**
4. The _____ noble gases that cannot bond to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form a bond, each has negative the electron configuration of _____ noble gas.
6. The atoms that make up _____ are ion bonded together by ionic bonds.
7. The _____ method can be used to write the formula for an ionic compound.
8. The _____ empirical formula is a compound denotes the ratio of ions in the compound. If the element is iron, write "Fe" if it is below. Change the subscript until an equals is made. If iron, write your answer on the line.

_____ **losing**

_____ **iron**

_____ **iron**

_____ **polyatomic**

_____ **cations**

_____ **iron (II) oxide**

9. Calcium boride is chemically written as _____ **gaining** two electrons.

10. A cation has a _____ **positive** charge.

11. A binary ionic compound contains only _____ **one** kind of cation and one kind of anion.

12. Substances that consist of more than one atom.

13. The Lewis dot diagram for chlorine, a group 7A element, has six _____ **electrons**.

14. A great deal of energy is _____ **produced** when an ionic compound is formed.

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