

Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility**: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.
saturated solution: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution**: a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible**: describes liquids that dissolve in each other. **immiscible**: describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47. 5.55×10^2 g AgNO_3
48. Solubility increases with pressure.
49. a. 1.6×10^{-2} g/L
b. 4.7×10^{-2} g/L
50. *Dilute* and *concentrated* are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.
a. 1.3M KCl
b. 3.3×10^{-1} M MgCl_2
52. 2.00×10^3 mL
53. a. 5.0×10^{-3} mol NaCl, 29 g NaCl
b. 1.0 mol KNO_3 , 1.0×10^{-2} g KNO_3
c. 2.5×10^3 mol CaCl_2 , 2.8 g CaCl_2
54. a. 2.3×10^1 g NaCl
b. 2.0 g MgCl_2
55. a. 16% (v/v) ethanol
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water
b. 1.5M KNO_3
c. 0.100M MgCl_2
58. The effective molality of the $\text{Ca}(\text{NO}_3)_2$ solution is 3m. The effective molality of the NaNO_3 solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g H_2O to 32.0 g CH_3OH .
63. a. 100.26°C
b. 101.54°C
64. a. -4.46°C
b. -2.2°C
65. a. -1.1°C
b. -0.74°C
c. -1.5°C

Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

World Of Chemistry Chapter 16 Test

William H. Brown, Thomas Poon



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Laboratory Organization. Automation D.J. Vonderschmitt, 2011-12-07 No detailed description available for Laboratory Organization Automation

Forensics For Dummies Douglas P. Lyle, 2019-05-07 Understand the real life science behind crime scene investigation Forensics For Dummies takes you inside the world of crime scene investigation to give you the low down on this exciting field Written by a doctor and former Law you ll learn what goes on at the morgue as you determine cause of death and you ll climb into the mind of a killer as you learn how forensic psychologists narrow down the suspect list Crime shows are entertaining but the reality is that most forensics cases aren't

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Combustion Chemistry and the Carbon Neutral Future Kenneth Brezinsky, 2023-02-16 As the demands for cleaner more efficient reduced and zero carbon emitting transportation increase the traditional focus of Combustion Chemistry research is stretching and adapting to help provide solutions to these contemporary issues Combustion Chemistry and the Carbon Neutral Future What will the Next 25 Years of Research Require presents a guide to current research in the field and an exploration of possible future steps as we move towards cleaner greener and reduced carbon combustion chemistry Beginning with a discussion of engine emissions and soot the book goes on to discuss a range of alternative fuels including hydrogen ammonia small alcohols and other bio oxygenates natural gas syngas and synthesized hydrocarbon fuels Methods for predicting and improving efficiency and sustainability such as low temperature and catalytic combustion chemical looping supercritical fluid combustion and diagnostic monitoring even at high pressure are then explored Some novel aspects of biomass derived aviation fuels and combustion synthesis are also covered Combining the knowledge and experience of an interdisciplinary team of experts in the field Combustion Chemistry and the Carbon Neutral Future What will the Next 25 Years of Research Require is an insightful guide to current and future focus areas for combustion chemistry researchers in line with the transition to greener cleaner technologies Provides insight on current developments in combustion chemistry as a tool for supporting a reduced carbon future Reviews

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