

Heat of Formation Worksheet

Use a standard enthalpies of formation table to determine the change in enthalpy for each of these reactions.



$$[(-411.0) + (-241.8)] - [(-426.7) + (-92.3)] = -133.8 \text{ kJ}$$



$$[2(-393.5)] - [2(-110.5)] = -566 \text{ kJ}$$



$$[(-393.5) + 2(-285.8)] - [(-74)] = -891.1 \text{ kJ}$$



$$[2(-285.8) + 2(-296.1)] - [2(-20.1)] = -1123.6 \text{ kJ}$$



$$[2(+33.9)] - [2(+90.4)] = -113 \text{ kJ}$$

Compound	ΔH_f (kJ/mol)	Compound	ΔH_f (kJ/mol)
$\text{CH}_4\text{(g)}$	-74.8	HCl(g)	-92.3
$\text{CO}_2\text{(g)}$	-393.5	$\text{H}_2\text{O(g)}$	-241.8
NaCl(s)	-411.0	$\text{SO}_2\text{(g)}$	-296.1
$\text{H}_2\text{O(l)}$	-285.8	$\text{NH}_4\text{Cl(s)}$	-315.4
$\text{H}_2\text{S(g)}$	-20.1	NO(g)	+90.4
$\text{H}_2\text{SO}_4\text{(l)}$	-811.3	$\text{NO}_2\text{(g)}$	+33.9
$\text{MgSO}_4\text{(s)}$	-1278.2	$\text{SnCl}_4\text{(l)}$	-545.2
MnO(s)	-384.9	SnO(s)	-286.2
$\text{MnO}_2\text{(s)}$	-519.7	$\text{SnO}_2\text{(s)}$	-580.7
NaCl(s)	-411.0	$\text{SO}_3\text{(g)}$	-296.1
NaF(s)	-569.0	$\text{SO}_3\text{(g)}$	-395.2
NaOH(s)	-426.7	ZnO(s)	-348.0
$\text{NH}_3\text{(g)}$	-46.2	ZnS(s)	-202.9

Standard Heat Of Formation Answer Key

Michael B. Mcelroy



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Table of Contents Standard Heat Of Formation Answer Key

1. Understanding the eBook Standard Heat Of Formation Answer Key
 - The Rise of Digital Reading Standard Heat Of Formation Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Standard Heat Of Formation Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Standard Heat Of Formation Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Standard Heat Of Formation Answer Key
 - Personalized Recommendations

- Standard Heat Of Formation Answer Key User Reviews and Ratings
- Standard Heat Of Formation Answer Key and Bestseller Lists
- 5. Accessing Standard Heat Of Formation Answer Key Free and Paid eBooks
 - Standard Heat Of Formation Answer Key Public Domain eBooks
 - Standard Heat Of Formation Answer Key eBook Subscription Services
 - Standard Heat Of Formation Answer Key Budget-Friendly Options
- 6. Navigating Standard Heat Of Formation Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Standard Heat Of Formation Answer Key Compatibility with Devices
 - Standard Heat Of Formation Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Standard Heat Of Formation Answer Key
 - Highlighting and Note-Taking Standard Heat Of Formation Answer Key
 - Interactive Elements Standard Heat Of Formation Answer Key
- 8. Staying Engaged with Standard Heat Of Formation Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Standard Heat Of Formation Answer Key
- 9. Balancing eBooks and Physical Books Standard Heat Of Formation Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Standard Heat Of Formation Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Standard Heat Of Formation Answer Key
 - Setting Reading Goals Standard Heat Of Formation Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Standard Heat Of Formation Answer Key
 - Fact-Checking eBook Content of Standard Heat Of Formation Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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