

NAME: _____

CLASS: _____

DATE: _____

1. What is the percent by mass of oxygen in MgO?

- | | | | |
|-------------------------|-----|-------------------------|-----|
| ☐ A | 50% | ☐ B | 60% |
| ☐ C | 20% | ☐ D | 40% |

2. How many molecules are in 2.5 mol of NaCl?

- | | | | |
|-------------------------|----------------|-------------------------|---------------------------------|
| ☐ A | 4.15 molecules | ☐ B | 1.51×10^{24} molecules |
| ☐ C | 1.50 molecules | ☐ D | 1.50×10^{24} molecules |

3. A student calculated the number of moles of methane (CH_4) present in a 100g sample of a gaseous mixture. If the student determined that there's 0.6 moles of methane present in the sample, what is the percent composition of methane in the sample?

- | | | | |
|-------------------------|------|-------------------------|-------|
| ☐ A | 9.6% | ☐ B | .096% |
| ☐ C | .9% | ☐ D | 90% |

4. A 15.67 g sample of a hydrate of magnesium carbonate was heated to drive off the water. The mass was reduced to 7.58 g. What is the formula of the hydrate?

- | | | | |
|-------------------------|---|-------------------------|---|
| ☐ A | $\text{MgCO}_3 \cdot 5\text{H}_2\text{O}$ | ☐ B | $\text{MgCO}_3 \cdot 4\text{H}_2\text{O}$ |
| ☐ C | $\text{MgCO}_3 \cdot 6\text{H}_2\text{O}$ | ☐ D | $\text{MgCO}_3 \cdot 1\text{H}_2\text{O}$ |

5. What is the empirical formula of a compound that is 81.82% carbon and 18.18% hydrogen?

- | | | | |
|-------------------------|---------------------------|-------------------------|------------------------|
| ☐ A | C_4H_{10} | ☐ B | C_2H_2 |
| ☐ C | CH_4 | ☐ D | C_3H_8 |

6. $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$

How many moles of oxygen are produced when 6.7 moles of KClO_3 decompose completely?

Star Test 10th Grade Chemistry

Thomas Griffiths



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