

<p>$4NaOH + 2F_2 \rightarrow 4NaF + 2H_2O + O_2$</p> <p>iii) (conc)</p> <p>+ Reaction with Cl_2, Br_2 and I_2 :</p> <p>$2NaOH + X_2 \rightarrow NaX + NaOX + H_2O$</p> <p>ii) (cold / dilute)</p> <p>iii) $6NaOH + 3X_2 \rightarrow 5NaX + NaXO_3 + 3H_2O$</p> <p>+ Sodium carbonate ($Na_2CO_3 \cdot 10H_2O$) - called washing soda.</p> <p>Anhydrous sodium carbonate is called soda ash.</p> <p>+ Sodium carbonate is prepared by</p> <p>(i) Le-Blanc process (ii) Ammonia-soda process or Solvay process (iii) Electrolytic process</p> <p>+ Le-Blanc process : The raw materials used are $NaCl$, conc. H_2SO_4, lime stone</p> <p>+ $Na_2CO_3 + CaS$ is called as black ash.</p> <p>+ Solvays process : The raw materials used are $NaCl$, NH_3 and $CaCO_3$.</p> <p>The by product in this process is $CaCl_2$.</p> <p>+ Sodium bi carbonate ($NaHCO_3$) is called as baking soda.</p> <p>+ $Na_2CO_3 + H_2O + CO_2 \rightarrow 2NaHCO_3$</p> <p>+ Aqueous solution of $NaHCO_3$ is alkaline due to anionic hydrolysis (HCO_3^-).</p> <p>+ Aqueous solution of $NaHCO_3$ cannot give any colour with phenolphthalein.</p> <p>+ The oxides of Be, Mg, Ca, Sr and Ba are basic therefore these elements are called alkaline earth metals.</p> <p>+ Radium is radio active element.</p>	<p>+ The outer electronic configuration is ns^2. They exhibit +2 oxidation state.</p> <p>+ The most abundant alkaline earth metal is calcium.</p> <p>+ Be and Mg do not give any flame colouration.</p> <p>+ Be shows diagonal relationship with Al.</p> <p>+ Alkaline earth metals form binary compounds with hydrogen called hydrides.</p> <p>The ionic nature of hydrides increases in the order</p> <p>$BeH_2 < MgH_2 < CaH_2 < SrH_2 < BaH_2$ (Ionic)</p> <p>+ Alkaline earth metals react with halogens to form their respective halides.</p> <p>+ Except Beryllium halides, all the other halides of alkaline earth metals are ionic.</p> <p>+ Alkaline earth metal fluorides are generally insoluble in water except BeF_2.</p> <p>+ Mg is used in the manufacture of alloys</p> <p>(i) Magnalium : $Al + Mg$ (85 – 90%) (1 – 15%)</p> <p>(ii) Elektron : $Mg + Zn$ (95 %) (5%)</p> <p>+ Gypsum on heating to $120 - 130^\circ C$ gives plaster of paris. ($2CaSO_4 \cdot H_2O$).</p> <p>+ Plaster of paris sets to hard mass in two stages</p> <p>$CaSO_4 \cdot \frac{1}{2}H_2O + \frac{3}{2}H_2O \xrightarrow{\text{setting stage}} CaSO_4 \cdot 2H_2O$</p> <p>$CaSO_4 \cdot 2H_2O \xrightarrow{\text{hardening stage}} CaSO_4 \cdot 2H_2O$ (ortho rhombic di hydrate) (Monoclinic dihydrate)</p> <p>+ Lime mortar is a mixture of 1 part of slaked lime, 3 parts of sand and water.</p> <p>+ Mortar mixed with cement is called cement mortar.</p> <p>+ Lime stone and clay when heated together give hydraulic mortar. This is used for bleaching purposes and as an antiseptic.</p>
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World Of Chemistry 2 1 Answers

**Don DeYoung, John Whitcomb, Donald
B. DeYoung**



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