

The number of moles of substances, when compared and reduced to the lowest terms, yields a mole ratio, as shown in the sample below. Indicate the mole ratio for each of the following equations:


2:1:2

1:1:1

4:3:2

1:2:1:1

2:1:2

2:1:1

Convert the following masses to moles by using the information on your periodic table.



$$34.0\text{g} \left(\frac{1\text{mol}}{17.0\text{g}} \right) = 2.00\text{mol}$$

2.00 mol


$$51.4\text{g} \left(\frac{1\text{mol}}{62.0\text{g}} \right) = 0.829\text{mol}$$

0.829 mol


$$4.50\text{g} \left(\frac{1\text{mol}}{98.1\text{g}} \right) = 0.0459\text{mol}$$

0.0459 mol


$$229\text{g} \left(\frac{1\text{mol}}{164.1\text{g}} \right) = 1.40\text{mol}$$

1.40 mol


$$79.5\text{g} \left(\frac{1\text{mol}}{35.0\text{g}} \right) = 2.27\text{mol}$$

2.27 mol

Convert the following to grams.



$$17.1\text{mol} \times \left(\frac{18.0\text{g}}{1\text{mol}} \right) = 308\text{g}$$

308 g.

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Julia Burdge



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