- Given the equation: C₁₂H₂₂O₁₁ + 12O₂ → 12CO₂ + 11H₂O + 5638 kJ
 - How much heat is released during the formation of 9.6 moles of CO₂?

b. How much heat is released during the formation of 0.036 moles of H₂O?

c. If 1026 grams of C₁₂H₂₂O₁₁ are consumed, how much heat is released?

d. If 23.76 grams of CO₂ are produced, how much heat is released?

Calculate the amount of heat (in Joules) required to warm 350.0 g of water from 30°C to 35°C. (Heat Capacity (C) for H2O is 4180 J/kg °C)

35.112 kJ of heat are added to a 500.0 gram sample of water initially at 7°C. Calculate the final temperature of the water sample. Be careful with units!

At= 35112 - 16.8°C

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