SINGLE REPLACEMENT REACTIONS WORKSHEET

1.	Using the activity series table answer the following questions:
	(a) list two metals that will react with Fe ²⁺ ions
	(b) list two metals that will lose electrons to Pb2+
	(c) list two ions that will take electrons from Zn
	(d) list two metals that will react with HCI
	(e) identify the best oxidizing agent (i.e. electron grabber)
2.	Using the activity series, how do you decide if a reaction will or will not take place between two species?
3.	Describe what you would see when a piece of Mg metal is placed in an aqueous solution of copper(II) sulfate. Writ the net ionic equation for the reaction that takes placed.
4.	What would happen to a silver dollar when placed in a 3 M HCI solution? Justify your answer.
5.	Given an ionic equation, how can you identify the spectator ions of a reaction?
6.	Do aluminum pop cans make a suitable containers for hydrochloric acid solutions? Write an equation that supports your answer.
7.	
8.	Rank order the following species in order of their increasing ability to lose electrons: Ni , Cu*, Cl, K, Mg
9.	What happens when a piece of calcium metal is placed in water? Write the molecular equation that supports your prediction.

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Jin-Ying Zhang

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