Six Types of Chemical Reaction Worksheet Answers

Balance the following reactions and indicate which of the six types of chemical reaction are being represented:

a) 2 NaBr + 1 Ca(OH)₂ → 1 CaBr₂ + 2 NaOH

Type of reaction: double displacement

b) 2 NH₃+ 1 H₂SO₄ → 1 (NH₄)₂SO₄

Type of reaction: synthesis

c) 4 C₈H₉O + 27 O₂ → 20 CO₂ + 18 H₂O

Type of reaction: combustion

d) 3 Pb + 2 H₃PO₄ → 3 H₂ + 1 Pb₃(PO₄)₂

Type of reaction: single displacement

e) 1 Li₂N + 3 NH₄NO₃ → 3 LiNO₃ + 1 (NH₄)₃N

Type of reaction: double displacement

f) 3 HBr + 1 Al(OH)₃ → 3 H₂O + 1 AlBr₃

Type of reaction: acid-base (double displacement)

g) Na₃PO₄ + 3 KOH → 3 NaOH + K₃PO₄

g) Najrou + 3 Non -> 3 Naon + Njrou

h) MgCl₂ + Li₂CO₃ → MgCO₃ + 2 LiCl

i) C₈H₁₂ + 9 O₂ → 6 CO₂ + 6 H₂O

i) Pb + FeSO₄ → PbSO₄ + Fe

k) CaCO₃ → CaO + CO₂

P₄ + 3 O₂ → 2 P₂O₃

m) 2 RbNO₃ + BeF₂ → Be(NO₃)₂ + 2 RbF

n) 2 AgNO₃ + Cu → Cu(NO₃)₂ + 2 Ag

O) C₃H₈O + 4 O₂ → 3 CO₂ + 3 H₂O

p) 2 CsHs + Fe → Fe(CsHs)₂

q) SeCl₀ + O₂ → SeO₂ + 3Cl₂

r) 2 Mgl₂ + Mn(SO₃)₂ → 2 MgSO₃ + Mnl₄

s) O₂ → O + O₂

t) 2 NO₂ → 2 O₂ + N₂

Reaction Type __double displacement____

Reaction Type double displacement

Reaction Type ____complete combustion____

Reaction Type ____single displacement____

Reaction Type ____decomposition_____

Reaction Type ___synthesis_____

Reaction Type ____double displacement____

Reaction Type ___single displacement____

Reaction Type ____complete combustion____

Reaction Type ___synthesis_____

Reaction Type ___single displacement____

Reaction Type ____double displacement____

Reaction Type ____decomposition_____

Reaction Type _____decomposition_____

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Thomas EJ Healy, Paul R Knight

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Table of Contents Skill Practice 28 Reactions Practice Answers

- 1. Understanding the eBook Skill Practice 28 Reactions Practice Answers
 - The Rise of Digital Reading Skill Practice 28 Reactions Practice Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Skill Practice 28 Reactions Practice Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skill Practice 28 Reactions Practice Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Skill Practice 28 Reactions Practice Answers
 - Personalized Recommendations
 - Skill Practice 28 Reactions Practice Answers User Reviews and Ratings
 - Skill Practice 28 Reactions Practice Answers and Bestseller Lists
- 5. Accessing Skill Practice 28 Reactions Practice Answers Free and Paid eBooks
 - Skill Practice 28 Reactions Practice Answers Public Domain eBooks
 - Skill Practice 28 Reactions Practice Answers eBook Subscription Services
 - Skill Practice 28 Reactions Practice Answers Budget-Friendly Options

- 6. Navigating Skill Practice 28 Reactions Practice Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Skill Practice 28 Reactions Practice Answers Compatibility with Devices
 - Skill Practice 28 Reactions Practice Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Skill Practice 28 Reactions Practice Answers
 - Highlighting and Note-Taking Skill Practice 28 Reactions Practice Answers
 - Interactive Elements Skill Practice 28 Reactions Practice Answers
- 8. Staying Engaged with Skill Practice 28 Reactions Practice Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skill Practice 28 Reactions Practice Answers
- 9. Balancing eBooks and Physical Books Skill Practice 28 Reactions Practice Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skill Practice 28 Reactions Practice Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Skill Practice 28 Reactions Practice Answers
 - Setting Reading Goals Skill Practice 28 Reactions Practice Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Skill Practice 28 Reactions Practice Answers
 - Fact-Checking eBook Content of Skill Practice 28 Reactions Practice Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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