

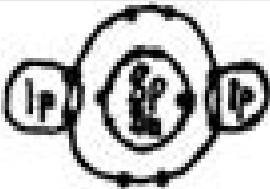
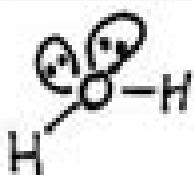
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Covalent Bonding Practice Worksheet

Explanation: Covalent compounds are compounds between nonmetal atoms. The atoms "start" with a number of electrons equal to their group number in the periodic table. Then they share electrons so that all the atoms have filled shells with a "magic number" of electrons. The magic number is 2 for hydrogen and 8 for C, N, O, F, Si, P, S, Cl, Se, Br and I. Remember that compounds made with a metal atom and a nonmetal atom make an ionic compound, not a covalent one, and compounds of metals are also different, these have metallic bonds and do not require a strict ratio.

Instructions: Fill in the blanks in the table. The first row is done for you; study the first row carefully before you start, especially the difference between different boxes (remember there are many ways to draw the same compound). Some rows have hints. If a box has an X you don't have to fill in that box.

Helpful Tips: 1) All atoms start with a number of outer shell electrons = their group number. You cannot change the element's group number. 2) If you have to choose between two shapes, the symmetrical, balanced one is usually best. 3) Try to figure the structures out yourself, don't cheat and use the Internet. ☺

Compound Formula	Full electronic structure	Lewis structure	Dot-cross diagram	Structural formula	3-d formula
H ₂ O		$\text{H} : \ddot{\text{O}} : \text{H}$	$\text{H}^{\times} \ddot{\text{O}}^{\times} \text{H}$	H-O-H	
NH ₃		$\begin{array}{c} \cdot\cdot \\ \text{H} : \text{N} : \text{H} \\ \\ \text{H} \end{array}$			
CH ₃ Br			$\begin{array}{c} \text{H} \\ \times \\ \text{H} : \text{C} : \text{H} \\ \times \\ : \text{Br} : \\ \cdot\cdot \end{array}$		

The Nature Of Covalent Bonding 16 1 Worksheet

Ian Pickup



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Table of Contents The Nature Of Covalent Bonding 16 1 Worksheet

1. Understanding the eBook The Nature Of Covalent Bonding 16 1 Worksheet
 - The Rise of Digital Reading The Nature Of Covalent Bonding 16 1 Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying The Nature Of Covalent Bonding 16 1 Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an The Nature Of Covalent Bonding 16 1 Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from The Nature Of Covalent Bonding 16 1 Worksheet
 - Personalized Recommendations
 - The Nature Of Covalent Bonding 16 1 Worksheet User Reviews and Ratings
 - The Nature Of Covalent Bonding 16 1 Worksheet and Bestseller Lists
5. Accessing The Nature Of Covalent Bonding 16 1 Worksheet Free and Paid eBooks
 - The Nature Of Covalent Bonding 16 1 Worksheet Public Domain eBooks
 - The Nature Of Covalent Bonding 16 1 Worksheet eBook Subscription Services
 - The Nature Of Covalent Bonding 16 1 Worksheet Budget-Friendly Options

6. Navigating The Nature Of Covalent Bonding 16 1 Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - The Nature Of Covalent Bonding 16 1 Worksheet Compatibility with Devices
 - The Nature Of Covalent Bonding 16 1 Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of The Nature Of Covalent Bonding 16 1 Worksheet
 - Highlighting and Note-Taking The Nature Of Covalent Bonding 16 1 Worksheet
 - Interactive Elements The Nature Of Covalent Bonding 16 1 Worksheet
8. Staying Engaged with The Nature Of Covalent Bonding 16 1 Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers The Nature Of Covalent Bonding 16 1 Worksheet
9. Balancing eBooks and Physical Books The Nature Of Covalent Bonding 16 1 Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection The Nature Of Covalent Bonding 16 1 Worksheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine The Nature Of Covalent Bonding 16 1 Worksheet
 - Setting Reading Goals The Nature Of Covalent Bonding 16 1 Worksheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of The Nature Of Covalent Bonding 16 1 Worksheet
 - Fact-Checking eBook Content of The Nature Of Covalent Bonding 16 1 Worksheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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