

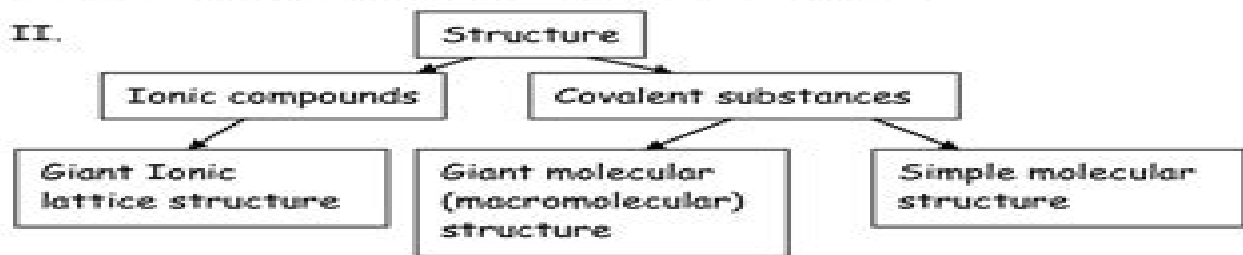
Answers to Worksheet 5 - Chemical Bonding
Science (Chemistry) Theory Workbook Pg 36 to 45

Check It (37 to 39)

1. C	2. B	3. D	4. A	5. B	6. A	7. D	8. B	9. B
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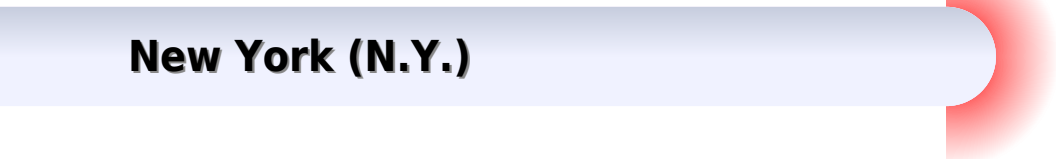
- 10a) calcium oxide and potassium chloride
 b) carbon dioxide and methane
 c) calcium oxide or potassium chloride (any one, an ionic one)
 d) carbon dioxide or methane (any one, a covalent one)
 e) helium (a noble gas , is monatomic, as single atoms)
- 11a) C :(2,4) : O :(2,6) : Na :(2,8,1) : Si :(2,8,4) : Cl :(2,8,7)
- b)(i) No. of neutrons = Nucleon no – no of protons = 9 – 4 = 5
 (ii) No. of neutrons = 28 – 14 = 14
- c) Beryllium and sodium (are metals, but, diamond does not conduct)
- d)(i) NaCl, sodium chloride
 (ii) Na₂O, sodium oxide
 (iii) CO₂, carbon dioxide or CO, carbon monoxide
 (iv) Na₂CO₃, sodium carbonate (C and O forms the radical ion CO₃²⁻)
- e) Carbon dioxide is a covalent compound with a simple molecular structure. It is made up of small molecules with weak intermolecular forces of attraction between molecules. Little amount of heat energy is needed to overcome these weak forces. Thus, it has a low boiling point and is a gas at r.t.p.

Notes: I. Type of Bonding - Ionic & Covalent



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1. Understanding the eBook Topic 6 Bonding Prentice Hall Textbook Answer
 - The Rise of Digital Reading Topic 6 Bonding Prentice Hall Textbook Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Topic 6 Bonding Prentice Hall Textbook Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Topic 6 Bonding Prentice Hall Textbook Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Topic 6 Bonding Prentice Hall Textbook Answer
 - Personalized Recommendations
 - Topic 6 Bonding Prentice Hall Textbook Answer User Reviews and Ratings
 - Topic 6 Bonding Prentice Hall Textbook Answer and Bestseller Lists
5. Accessing Topic 6 Bonding Prentice Hall Textbook Answer Free and Paid eBooks
 - Topic 6 Bonding Prentice Hall Textbook Answer Public Domain eBooks
 - Topic 6 Bonding Prentice Hall Textbook Answer eBook Subscription Services
 - Topic 6 Bonding Prentice Hall Textbook Answer Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Topic 6 Bonding Prentice Hall Textbook Answer Compatibility with Devices
 - Topic 6 Bonding Prentice Hall Textbook Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Topic 6 Bonding Prentice Hall Textbook Answer
 - Highlighting and Note-Taking Topic 6 Bonding Prentice Hall Textbook Answer
 - Interactive Elements Topic 6 Bonding Prentice Hall Textbook Answer
8. Staying Engaged with Topic 6 Bonding Prentice Hall Textbook Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Topic 6 Bonding Prentice Hall Textbook Answer
9. Balancing eBooks and Physical Books Topic 6 Bonding Prentice Hall Textbook Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Topic 6 Bonding Prentice Hall Textbook Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Topic 6 Bonding Prentice Hall Textbook Answer
 - Setting Reading Goals Topic 6 Bonding Prentice Hall Textbook Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Topic 6 Bonding Prentice Hall Textbook Answer
 - Fact-Checking eBook Content of Topic 6 Bonding Prentice Hall Textbook Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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