

- What is the proper symbol for an atom with 45 protons, 101 neutrons and 45 electrons?
 A) $^{146}_{45}\text{Rh}$
 B) $^{146}_{101}\text{Md}$
 C) $^{146}_{45}\text{Rh}$
 D) $^{146}_{101}\text{Fm}$
 E) I am not sure.
- What is the name of the molecular compound IF_7 ?
 A) Moniodide fluoride
 B) Moniodine heptafluoride
 C) Iodine heptafluoride
 D) Iodine fluoride
 E) I am not sure.
- Which element has the following electron configuration: $1s^2 2s^2 2p^4 3s^2 3p^4$?
 A) N
 B) S
 C) P
 D) Se
 E) I am not sure.
- Write the formula unit formed by copper (II) oxide.
 A) CuO
 B) Cu_2O
 C) CuO_2
 D) Cu_2O_2
 E) I am not sure.
- Balance the following reaction:
 $__\text{C}_2\text{H}_6(\text{g}) + __\text{O}_2(\text{g}) \rightarrow __\text{CO}_2(\text{g}) + __\text{H}_2\text{O}(\text{g})$
 A) 1, 3, 2, 1
 B) 2, 5, 2, 2
 C) 1, 3, 2, 2
 D) 2, 6, 4, 4
 E) I am not sure.
- Element X has three natural isotopes: X-4 (4.240 amu), X-5 (5.025 amu), and X-7 (7.349 amu). Calculate the mass of element x to two decimal places given the abundance X-4 is 20.6 % and X-5 is 36.7 %.
 A) 5.86 amu
 B) 5.61 amu
 C) 5.75 amu
 D) 5.65 amu
 E) I am not sure.
- What is the charge for the cobalt ion in CoO ?
 A) $1+$
 B) $2+$
 C) $2-$

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