

# Chemistry 105

## Chapter 3 Review Problems

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1. How many molecules are present in 180 grams of water? (The atomic weights of H and O are 1 g/mol and 16 g/mol, respectively.)  
a)  $6.0 \times 10^{24}$   
b) 10  
c)  $5.6 \times 10^{-23}$   
d)  $1.8 \times 10^7$   
e)  $180 \times 10^{23}$
2. The atomic mass of sulfur is given as 32.07 g/mol. What is the mass of an average atom of sulfur?  
a)  $5.33 \times 10^{-23}$  g  
b) 32.07 g  
c)  $32.07 \times 10^{-23}$  kg  
d)  $193 \times 10^{-22}$  g  
e)  $6.02 \times 10^{23}$  g
3. What is the mass of C present in 60 g of  $C_2H_6$ ?  
a) 12 g  
b) 24 g  
c) 48 g  
d) 60 g  
e) 6 g
4. What is the molar mass of  $C_2H_6O$ ?  
a) 24 g/mol  
b) 32 g/mol  
c) 7 g/mol  
d) 35 g/mol  
e) 46 g/mol
5. What is the correct percentage by mass of nitrogen in HCN?  
a) 33 %  
b) 14%  
c) 27%  
d) 52%  
e) 98%
6. What is the chief danger that is anticipated if  $CO_2$  levels get too high in the atmosphere?  
a) people will develop skin cancer from exposure to UV light  
b) the atmosphere will develop an unpleasant odor  
c) the temperature of the planet may climb by several degrees  
d) plant life will be exterminated  
e) The  $CO_2$  will remove ozone from the stratosphere

# Standardized Test Practice Chapter 3 Chemistry Answers

**The Princeton Review**



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