

4. A solution contains 2.0 g of NaCl in 100 mL of water. The solution is heated to 100°C, and the water is evaporated. How much NaCl is left? (The boiling point of water is 100°C.)
- 2.0 g
 - 10.0 g
 - 10.0 mL
 - 10.0 L
5. A solution contains 2.0 g of NaCl in 100 mL of water. The solution is heated to 100°C, and the water is evaporated. How much NaCl is left? (The boiling point of water is 100°C.)
- 2.0 g
 - 10.0 g
 - 10.0 mL
 - 10.0 L
6. A solution contains 2.0 g of NaCl in 100 mL of water. The solution is heated to 100°C, and the water is evaporated. How much NaCl is left? (The boiling point of water is 100°C.)
- 2.0 g
 - 10.0 g
 - 10.0 mL
 - 10.0 L
7. A solution contains 2.0 g of NaCl in 100 mL of water. The solution is heated to 100°C, and the water is evaporated. How much NaCl is left? (The boiling point of water is 100°C.)
- 2.0 g
 - 10.0 g
 - 10.0 mL
 - 10.0 L
8. A solution contains 2.0 g of NaCl in 100 mL of water. The solution is heated to 100°C, and the water is evaporated. How much NaCl is left? (The boiling point of water is 100°C.)
- 2.0 g
 - 10.0 g
 - 10.0 mL
 - 10.0 L
9. A solution contains 2.0 g of NaCl in 100 mL of water. The solution is heated to 100°C, and the water is evaporated. How much NaCl is left? (The boiling point of water is 100°C.)
- 2.0 g
 - 10.0 g
 - 10.0 mL
 - 10.0 L
10. A solution contains 2.0 g of NaCl in 100 mL of water. The solution is heated to 100°C, and the water is evaporated. How much NaCl is left? (The boiling point of water is 100°C.)
- 2.0 g
 - 10.0 g
 - 10.0 mL
 - 10.0 L

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N Colangelo



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Glencoe Science McGraw-Hill Staff, 2001-06

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