

NAME: \_\_\_\_\_

CLASS: \_\_\_\_\_

DATE: \_\_\_\_\_

1. What is the percent by mass of oxygen in MgO?

- |                         |     |                         |     |
|-------------------------|-----|-------------------------|-----|
| <input type="radio"/> A | 50% | <input type="radio"/> B | 60% |
| <input type="radio"/> C | 20% | <input type="radio"/> D | 40% |

2. How many molecules are in 2.5 mol of NaCl?

- |                         |                |                         |                                 |
|-------------------------|----------------|-------------------------|---------------------------------|
| <input type="radio"/> A | 4.15 molecules | <input type="radio"/> B | $1.51 \times 10^{24}$ molecules |
| <input type="radio"/> C | 1.50 molecules | <input type="radio"/> D | $1.50 \times 10^{24}$ molecules |

3. A student calculated the number of moles of methane ( $\text{CH}_4$ ) present in a 100g sample of a gaseous mixture. If the student determined that there's 0.6 moles of methane present in the sample, what is the percent composition of methane in the sample?

- |                         |      |                         |       |
|-------------------------|------|-------------------------|-------|
| <input type="radio"/> A | 9.6% | <input type="radio"/> B | .096% |
| <input type="radio"/> C | .9%  | <input type="radio"/> D | 90%   |

4. A 15.67 g sample of a hydrate of magnesium carbonate was heated to drive off the water. The mass was reduced to 7.58 g. What is the formula of the hydrate?

- |                         |   |                         |   |
|-------------------------|---|-------------------------|---|
| <input type="radio"/> A | $\text{MgCO}_3 \cdot 5\text{H}_2\text{O}$ | <input type="radio"/> B | $\text{MgCO}_3 \cdot 4\text{H}_2\text{O}$ |
| <input type="radio"/> C | $\text{MgCO}_3 \cdot 6\text{H}_2\text{O}$ | <input type="radio"/> D | $\text{MgCO}_3 \cdot 1\text{H}_2\text{O}$ |

5. What is the empirical formula of a compound that is 81.82% carbon and 18.18% hydrogen?

- |                         |                           |                         |                        |
|-------------------------|---------------------------|-------------------------|------------------------|
| <input type="radio"/> A | $\text{C}_4\text{H}_{10}$ | <input type="radio"/> B | $\text{C}_2\text{H}_2$ |
| <input type="radio"/> C | $\text{CH}_4$             | <input type="radio"/> D | $\text{C}_3\text{H}_8$ |

6.  $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$

How many moles of oxygen are produced when 6.7 moles of  $\text{KClO}_3$  decompose completely?

# Star Test 10th Grade Chemistry

**SJ Ball**



## Star Test 10th Grade Chemistry:

*Lubricant Additives* Leslie R. Rudnick, 2017-07-12 This indispensable book describes lubricant additives their synthesis chemistry and mode of action All important areas of application are covered detailing which lubricants are needed for a particular application Laboratory and field performance data for each application is provided and the design of cost effective environmentally friendly technologies is fully explored This edition includes new chapters on chlorohydrocarbons foaming chemistry and physics antifoams for nonaqueous lubricants hydrogenated styrene diene viscosity modifiers alkylated aromatics and the impact of REACH and GHS on the lubricant industry     Improving High School Paul Warren, 2005

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