

States Of Consciousness

AP Psychology

Consciousness

- **Dual Processing:** The principle that information is perceived, remembered and thought both consciously and unconsciously at the same time.
 - **Conscious Processing:** Deliberate ("High Road")
 - **Unconscious Processing:** Automatic ("Low Road")
- **Blindsight:** The ability for a person who cannot see to be able to respond to visual stimulus (like if a blind person was able to avoid a trash can in front of him)
- **Selective Attention:** Focusing on one particular stimulus out of all the stimulus around oneself.
- **Cocktail Party Effect:** The ability to attend to only one voice out of many.
- **Inattentional Blindness:** The inability to notice a visible object because your attention is elsewhere.
- **Change Blindness:** When there is a visual stimulus alters but the observer is unaware of the alteration.
- **Change Deafness:** When an audible stimulus (noise) alters, including its tone, pitch, and/or volume but the listener is unaware of the alteration.

Sleep

- **Circadian Rhythm:** The body's internal (biological) sleep-wake clock.
 - It can be affected by age and experience
- **Stages of Sleep**
 - **Hypnagogic State:** The dreamlike state which occurs before sleep.
 - **Quiet Sleep:** Occurs stages 1 to 4, nRem (not Rem).
 - **Active Sleep:** Rem, dreams most likely occur in active sleep.
 - **Individual Stages of Sleep**
 - **Stage 1**
 - Alpha waves
 - Signalled by twitches
 - Lasts 10minutes
 - If awoken people in stage 1 don't think they were asleep
 - **Stage 2**
 - Spindles
 - Stage 2 takes up 50% of a person's sleep time
 - **Stage 3**
 - Delta waves
 - Lasts about 30minutes
 - The body begins slowing down (example: breathing slows)

States Of Consciousness Study Guide Answers

David S. Moore



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subsequent Δh values
 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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web chemistry 120 hess s law worksheet calculate ΔH for the reaction $C_2H_4(g) + 2Cl_2(g) \rightarrow C_2H_6(g) + 2HCl(g)$ from the following data $C_2H_4(g) + 3O_2(g) \rightarrow 2CO_2(g) + 2H_2O(l)$ $\Delta H = -1411 \text{ kJ/mol}$ $C_2H_6(g) + 7/2 O_2(g) \rightarrow 2CO_2(g) + 3H_2O(l)$ $\Delta H = -1560 \text{ kJ/mol}$ $H_2(g) + 1/2 O_2(g) \rightarrow H_2O(l)$ $\Delta H = -285 \text{ kJ/mol}$

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web libretexts states hess s law states that regardless of the multiple stages or steps of a reaction the total enthalpy change for the reaction is the sum of all changes libretexts *quot hess s law quot* 2019 which determines the enthalpy change value hess s law is written as $\Delta H_{\text{target}} = \Delta H_1 + \Delta H_2 + \Delta H_3$ or $\Delta H_{\text{target}} = \sum \Delta H_{\text{known}}$

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web hess s law states that if a process can be expressed as the sum of two or more steps the enthalpy change for the overall process is the sum of the ΔH values for each step to use hess s law two principles must be understood one if an equation is reversed the sign of the ΔH value is also reversed two if an equation is multiplied by a

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web hess s law states that the total enthalpy change for the reaction will be the sum of all those changes no matter how many different steps or stages in the reaction there are cohen 2016 the equations for the reactions in the experiment done are as follows $1. NaOH(s) \rightarrow Na^+(aq) + OH^-(aq)$

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web this activity provides a demonstration of hess law using three reactions the solubility naoh in water the solubility naoh in hcl and the reaction of a solution of hcl and a solution of naoh

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web hess s law worksheet answers 1 calculate ΔH for the reaction $C_2H_4(g) + H_2(g) \rightarrow C_2H_6(g)$ from the following data $C_2H_4(g) + 3 O_2(g) \rightarrow 2 CO_2(g) + 2 H_2O(l) \quad \Delta H = -1411 \text{ kJ}$ $C_2H_6(g) + \frac{7}{2} O_2(g) \rightarrow 2 CO_2(g) + 3 H_2O(l) \quad \Delta H = -1560 \text{ kJ}$

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web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kJ in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and ΔH values to determine the heat of reaction for the reaction C_2H_4

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of $NaOH$ combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web hess s law enthalpy of reaction and hess s law abstract the goal of this lab is to perform experiments and practice hess s law by determining enthalpy change in chemical reactions in this lab three experiments were conducted in order to test hess s law the experiments required observing and collecting data on three different reactions

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{aq} + \text{OH}^- \text{aq}$ $\Delta H = -44.5 \text{ kJ/mol}$ 2 $\text{NaOH (s)} + \text{H}^+ \text{aq} \rightarrow \text{Na}^+ \text{aq} + \text{H}_2\text{O (l)}$ $\Delta H = -57.3 \text{ kJ/mol}$ 3