

A1 States of matter

1. a) i) Ammonium chloride (1)
 ii) Diffusion (1)
 Diffusion is the movement of particles from an area of higher concentration to an area of lower concentration until the particles are evenly distributed. (2)
 iii) The ammonia solution gave off ammonia gas and the hydrochloric acid gave off hydrogen chloride gas. The particles of the two gases diffused along the tube, collided and reacted to form ammonium chloride. (2)



- b) i) (1)
 ii) The distilled water had a higher water content than the cytoplasm inside the paw paw cells and the cell membranes were differentially permeable. The water molecules were able to move through the membranes into the cells causing the cells to swell and the strips to increase in volume. (3)
 iii) Osmosis (1)
 c) Sodium chloride draws water out of the cells of the fish and any micro-organisms by osmosis. Water is, therefore, not available for chemical reactions in the cells, some of which cause decay and micro-organisms can't grow. (2)

2. a)

| Property | Solid | Liquid | Gas |
|--------------------------|---|---------------------------------------|--|
| Volume | | definite | variable; the volume is the same as the entire container |
| Arrangement of particles | packed closely together, usually in a regular way | | randomly arranged with large spaces between |
| Energy of particles | have very small amounts of kinetic energy | have medium amounts of kinetic energy | |

- b) i) The particles have large spaces between them, so they can be easily pushed closer together. (1)
 ii) The particles are packed tightly together with very little empty space between. (1)
 iii) The particles move around rapidly and have weak forces of attraction between them, so they spread out to fill any available space in the container. (2)

- c) i) A: Melting B: Boiling or evaporation (4)
 C: Freezing D: Condensation (4)
 ii) Iodine or carbon dioxide or ammonium chloride or naphthalene (1)
 d) i) Liquid (1)
 ii) 54 °C (1)
 iii) (1)



A2 Mixtures and separations

1. a) i)

| | Pure substance | Mixture |
|-------------|--------------------|---|
| Composition | | variable |
| Properties | fixed and constant | variable; the components retain their individual properties |

(3)

- ii) An element is a pure substance that cannot be broken down into any simpler substances by any ordinary chemical or physical means. A compound is a pure substance that contains two or more different types of element that are bonded together chemically in fixed proportions and in such a way that their properties have changed. (2)
 b) i) The particles in a suspension are larger than those in a colloid. (1)
 ii) The particles in a suspension settle if left undisturbed, whereas the particles in a colloid never settle. (1)
 iii) Light does not pass through a suspension, whereas most colloids scatter light. (1)
 c) i) A solution is a homogeneous mixture consisting of two or more components, one of which is usually a liquid. (1)
 ii) Solubility is the mass of solute that will saturate 100 g of solvent at a given temperature. (1)
 d) i) 17 g per 100 g water (1)
 ii) 45 °C (1)
 iii) At 76 °C, 54 g of Z saturate 100 g water
 At 36 °C, 12.5 g of Z saturate 100 g water
 \therefore mass of Z crystallising out of a saturated solution containing 100 g water = $54 - 12.5 \text{ g} = 41.5 \text{ g}$ (3)
 iv) At 62 °C, 39 g of Z saturate 100 g water
 \therefore at 62 °C, 45 g of Z saturate $\frac{100}{39} \times 45 \text{ g water} = 115.4 \text{ g of water}$ (2)
 v) At 55 °C, 33 g of Z saturate 100 g water
 \therefore at 55 °C, $\frac{33}{100} \times 350 \text{ g of Z saturate } 350 \text{ g water} = 115.5 \text{ g of Z}$ (2)

States Of Matter Chemistry Workbook Answers

Arshad Iqbal



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