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Slk 200 Kompressor Repair Manual

**A. K. Legg, Haynes Publishing
PLC, Spencer Drayton**



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{O}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g})$ $\Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + \text{O}_2(\text{g})$ $\Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + 3 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ $\Delta H = 283 \text{ kJ}$

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302 g co₂ g 2 h₂ o l h 1411 kJ mole c 2 h 6 g 7 2 o 2 g 2 co₂ g 3 h₂ o l h 1560 kJ mole h 2 g 1 2 o 2 g h₂ o l h 285 kJ mole

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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