

8.2 The Nature of Covalent Bonding

Essential Understanding

Covalent bonds form when atoms share electrons.

Lesson Summary

The Octet Rule in Covalent Bonding

Covalent compounds are most stable when each atom has eight electrons.

- ▶ Single, double, and triple covalent bonds depend on the number of pairs of electrons shared between two atoms.
- ▶ Atoms form double or triple covalent bonds if they can attain a noble gas structure by doing so.

Type of Covalent Bond	Attributes
Single	One shared electron pair with one electron from each atom
Double	Two shared electron pairs with two electrons from each atom
Triple	Three shared electron pairs with three electrons from each atom

Coordinate Covalent Bonds In a coordinate covalent bond, one atom contributes both electrons in the bonding pair.

- ▶ One atom may contribute a pair of unshared electrons to a bond to give both atoms an inert gas configuration.
- ▶ Coordinate covalent bonds can also occur in polyatomic ions, such as NH_4^+ .

Exceptions to the Octet Rule Some molecules have fewer, or more, than a complete octet of valence electrons.

- ▶ Molecules that have an odd number of total valence electrons cannot satisfy the octet rule.
- ▶ Some molecules that have an even number of valence electrons may also fail to follow the octet rule.

The Nature Of Covalent Bonding 16 Worksheet

VM Jensen



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