

Comments on specific questions

**Question 7**

A large number of candidates gave the correct answer of **B** but many selected **D**. This shows the importance of reading the question carefully. Whilst statement **D** is clearly a correct fact, it does not show that "graphite and diamond are different forms of the same element".

**Question 8**

In this case many candidates gave the correct response, **C**. However a significant number of candidates gave each of the other three options in almost equal number, presumably having mis-counted the protons and then guessed at one of the other responses.

**Question 10**

Whilst many candidates gave the correct answer, **A**, a significant minority gave **D**. It would appear that candidates in the latter group considered the information that ammonia (an alkali) reacts with carbon dioxide (an acid) and therefore were looking for the formation of a salt. However this conclusion was not justified since the formula of this product did not correspond to the atoms that were specified as being present.

**Question 16**

This question proved challenging for many candidates. However stronger candidates gave the correct answer, **D**, since the bonding within both oil and water molecules is the same – covalent – making **D** the only incorrect statement. Many candidates gave **C** as their response. This is a correct statement as, although overall oil is less dense than water, in a mixture such as oil, there will be molecules with a large variety of relative molecular masses.

**Question 20**

This question was only answered correctly by the stronger candidates. Both **B** and **C** were chosen more frequently. To reach a valid conclusion candidates had to realise that, from the stoichiometry of the equation for one mole of each carbonate, the volume of carbon dioxide released would be the same but, that if insoluble sulfates were being formed, the reaction would not go to completion.

**Question 30**

A minority of candidates recognised that hydrated sodium carbonate loses water of crystallisation on heating to give the correct response, **B**. A large number of candidates, however, thought that the anhydrous compound further decomposed to release carbon dioxide. This led to **D** as an answer.

**Question 34**

This question was the most challenging on the paper. To reach the conclusion it was necessary to note, in the stem of the question, that the molecule had to be both unsaturated and a hydrocarbon. It is probable that the candidates who gave **B** as an answer did not notice that the first structure is not a hydrocarbon while those who gave **C** as answer merely looked for the presence of any double bond.

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**Terry C. Jones**



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