

## Worksheet: Mole Calculations & Percent Composition

A. Mole Calculations. You will need a periodic table to look up molecular weights. Do each type of calculation until you feel that you could do it without looking at notes. If you are very comfortable with a problem, skip down to harder ones. The answers will be on the web.

1. \_\_\_\_\_ How many moles of Na are in 42 g of Na?
2. \_\_\_\_\_ How many moles of O are in 8.25 g of O?
3. \_\_\_\_\_ How much does 2.18 mol of Cu weigh?
4. \_\_\_\_\_ What is the mass of 0.28 mol of iron?
5. \_\_\_\_\_ How many atoms are in 7.2 mol of chlorine?
6. \_\_\_\_\_ How many atoms are in 36 g of bromine?
7. \_\_\_\_\_ How many moles are in  $1.0 \times 10^7$  atoms?
8. \_\_\_\_\_ What is the mass of  $1.20 \times 10^{25}$  atoms of sulfur?
9. \_\_\_\_\_ How many moles of CO molecules are in 52 g of CO?
10. \_\_\_\_\_ How many moles of  $C_2H_6$  are in 124 g?
11. \_\_\_\_\_ How many moles of  $CCl_4$  are there in 56 g?
12. \_\_\_\_\_ How much does 2.50 mol of  $H_2SO_4$  weigh?
13. \_\_\_\_\_ How much does 0.25 mol of  $Fe_2O_3$  weigh?
14. \_\_\_\_\_ How many molecules are there in 52 g of CO?
15. \_\_\_\_\_ How many formula units are in 22.4 g  $SnO_2$ ?
16. \_\_\_\_\_ How many molecules are in 116 g  $CCl_4$ ?
17. \_\_\_\_\_ What is the mass of  $3.01 \times 10^{23}$  formula units of  $Fe_2O_3$ ?
18. \_\_\_\_\_ What is the mass of  $1.2 \times 10^{25}$  molecules of CO?
19. \_\_\_\_\_ How many O atoms are in 1.25 mol of  $SO_2$ ?
20. \_\_\_\_\_ How many moles of O atoms do you have when you have  $1.20 \times 10^{25}$   $N_2O_5$  molecules?
21. \_\_\_\_\_ How many formula units are in 5.33 mol of  $CuCl_2$ ?
22. \_\_\_\_\_ How many copper atoms are in 5.33 mol of  $CuCl_2$ ?
23. \_\_\_\_\_ How many moles of Cl atoms are in 5.33 mol of  $CuCl_2$ ?

# The Mole The Central Unit Of Chemistry Worksheet

**Tina Hirsch**



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