

8.2**THE NATURE OF COVALENT BONDING****Section Review****Objectives**

- State a rule that usually tells how many electrons are shared to form a covalent bond.
- Describe how electrons are formulated.
- Predict when two atoms are likely to be joined by a double or a triple covalent bond.
- Distinguish between a single covalent bond and other covalent bonds.
- Describe how the strength of a covalent bond is related to its bond dissociation energy.
- Describe how resonance structures explain bonding.
- Identify some exceptions to the octet rule.

Vocabulary

- | | | |
|------------------------|----------------------------|----------------------------|
| • single covalent bond | • double covalent bonds | • polyatomic ion |
| • structural formulas | • triple covalent bonds | • bond dissociation energy |
| • unshared pairs | • coordinate covalent bond | • resonance structures |

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are presented in this section. Each blank can be completed with a term, short phrase, or number.

- When atoms share electrons to gain the ____1____ configuration of a noble gas, the bonds formed are ____2____. A ____3____ pair of valence electrons constitutes a ____4____ covalent bond. Pairs of valence electrons that are not shared between atoms are called ____5____. Sometimes two or three pairs of electrons may be shared to give ____6____ covalent bonds. In some cases, only one of the atoms in a bond provides the pair of bonding electrons; this is a ____7____ ____8____. It is required to break covalent bonds between atoms. The total energy required to break the bond between two covalently bonded atoms is known as the ____9____.
- When it is possible to write two or more valid electron dot formulas for a molecule or ion, each formula is referred to as a ____10____.

The Nature Of Covalent Bonding Answer Key

O García



The Nature Of Covalent Bonding Answer Key:

Reviewing **The Nature Of Covalent Bonding Answer Key**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**The Nature Of Covalent Bonding Answer Key**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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