

Study Question 2a

New Chemical Reaction

Name: Arden E. BennettDate: 1/22/2021

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Chemical Thermodynamic Relationships

1. Standard reaction has worked the following reactions:

(a) $\text{Acetylene} + \text{hydrogen} \rightarrow \text{ethane}$ with standard enthalpy of formation: $\Delta H_f^\circ(\text{ethane}) = -84.68 \text{ kJ/mol}$



(b) $\text{C}_2\text{H}_2(\text{g}) + \text{hydrogen} \rightarrow \text{ethyne}$ with standard enthalpy of formation: $\Delta H_f^\circ(\text{ethyne}) = 107.8 \text{ kJ/mol}$



2. Standard free energy values:

(a) $\Delta G_f^\circ(\text{C}_2\text{H}_6, \text{g}) = -32.81 \text{ kJ/mol}$ (b) $\Delta G_f^\circ(\text{C}_2\text{H}_2, \text{g}) = 69.89 \text{ kJ/mol}$

(c) $\Delta G_f^\circ(\text{H}_2, \text{g}) = 0 \text{ kJ/mol}$ (d) $\Delta G_f^\circ(\text{H}_2\text{O}, \text{g}) = -228.57 \text{ kJ/mol}$

Combustion Reaction

1. Standard combustion reactions for the following compounds:

(a) $\text{C}_2\text{H}_6(\text{g}) + 7/2\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O}(\text{g}) + 1560 \text{ kJ}$

(b) $\text{C}_2\text{H}_2(\text{g}) + 5/2\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$

Equilibrium and Thermodynamic Relationships

1. Write an equation for the decomposition of ammonium carbonate:



2. Write an equation for the reaction of carbon dioxide with an alkoxide:



Source: <https://www.ck12.org/chemistry/thermodynamics-and-thermochemical-equations/>

Skill Practice 36 Answers

Alan Parker, Jan Faulkner



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