

**Solutions** are homogeneous mixtures of two or more substances. The substance present in smaller amount in the mixture is called the **solute** while the substance that has greater amount is the **solvent**. The solute is dissolved in the solvent. Homogenous means that you cannot identify the particles of the solute and the solvent separately in the solution. Mixture means that the solute and the solvent retain their individual properties and that the solute and the solvent are only physically combined. A solution can be gaseous like air, liquid like soda or solid like metal brass (alloy of copper and zinc).

**Aqueous solutions** are solutions in which the solvent is water.

**Electrolytic solutions** are solutions that can conduct electricity due to the presence of dissolved ions.

**Non-electrolytic solutions** are solutions that cannot conduct electricity because the solutes do not ions when dissolved in water.

**Electrolyte** is a solute that results to an electrolytic solution when dissolved in water.

**Non-electrolyte** is a solute that does not result to an electrolytic solution when dissolved in water.

In strong electrolytic solutions (solutions of strong electrolytes), the electrolyte is assumed to be 100% dissociated. **Dissociation** pertains to the separation of positive and negative ions in an electrolyte. If we have a strong electrolyte like Potassium Nitrate ( $\text{KNO}_3$ ), what this means is that  $\text{KNO}_3$  completely separates to  $\text{K}^+$  and  $\text{NO}_3^-$  when dissolved in water. No  $\text{KNO}_3$  is left undissociated.

The table below shows strong and weak electrolytes as well as non-electrolytes.

| Strong electrolytes      | Weak electrolytes        | Non-electrolytes                          |
|--------------------------|--------------------------|-------------------------------------------|
| HCl                      | $\text{CH}_3\text{COOH}$ | $(\text{NH}_2)_2\text{CO}$                |
| $\text{HNO}_3$           | HF                       | $\text{CH}_3\text{OH}$                    |
| $\text{HClO}_4$          | $\text{HNO}_2$           | $\text{C}_2\text{H}_5\text{OH}$           |
| $\text{H}_2\text{SO}_4$  | $\text{NH}_3$            | $\text{C}_6\text{H}_{12}\text{O}_6$       |
| NaOH                     | $\text{H}_2\text{O}$     | $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ |
| $\text{Ba}(\text{OH})_2$ |                          |                                           |
| Ionic compounds          |                          |                                           |

Therefore, NaOH completely dissociates to form  $\text{Na}^+$  and  $\text{OH}^-$



The (aq) subscript indicates that the ions are dissolved in water. Notice that the arrow is unidirectional which means that the reverse reaction does not happen spontaneously.

For weak electrolytes like HF, the process is reversible and eventually reaches an equilibrium state in which the rate with which HF dissociate to  $\text{H}^+$  and  $\text{F}^-$  ions is equal to the rate with which  $\text{H}^+$  and  $\text{F}^-$  ions recombine to form HF.



### Precipitation Reactions

Precipitation reactions are reactions that lead to the formation of precipitates which are water insoluble, solid substances.

# Study Guide Reactions In Aqueous Solutions Answers

**Barron's Educational Series, Mark  
Kernion, Joseph A. Mascetta**



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