

Unit 3 - Periodicity

Trends in physical properties

Effective Nuclear Charge

- The nuclear charge of an atom is given by its atomic number. (e.g. Oxygen is +8)
- The outer electrons of the atom **do not experience the full attraction** of the nuclear charge because they are **shielded from the nucleus because of the repulsion by the inner electrons** and so the effective nuclear charge is lesser than the nuclear charge.
- Nuclear charge increases across a period, but remains constant down a group

Atomic Radius

- the distance from the nucleus to the outermost electrons of an atom.
- Atomic radius decreases across a period, as attraction between the nucleus and outermost electrons increases, and increases down a group, as the number of shells increases

Ionic Radius

- Cations: As there are less electrons than protons, the outer shell is lost and the ionic radius of a cation is lesser than the atomic radius of the atom.
 - e.g. Na has a atomic radius of 186×10^{-12} while Na^+ has a ionic radius of 98×10^{-12}
- Anions: As there are more electrons than protons, there is more electron repulsion and less attraction, so the ionic radius is more than the atomic radius of the atom.
 - e.g. Cl has an atomic radius of 99×10^{-12} , but Cl^- has an ionic radius of 181×10^{-12}

Ionization Energy

- The first ionization energy of an element is the energy required to remove one mole of electrons from one mole of gaseous atoms in their ground state.
- Increases across a period as atomic bonds are stronger, and decreases down a group as there are more outshells and attraction is weaker.

Exceptions

- Groups 2 and 3
 - Elements of group 3 have a lower lower ionization energy than elements of Group 2.
 - Elements of Group 3 have the electron configuration of s^2p^1 , which means that the electron removed is from the p orbital, which is unpaired, making it easier to remove.
 - However, elements of Group 2 have an electron configuration of s^2 , which is a complete orbital and when an electron is removed, it is removed from this orbital, and it is relatively harder to remove the electron.
- Groups 15 and 16
 - Elements of Group 16 have a lower ionization energy compared to elements of Group 15.

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G Psacharopoulos



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Table of Contents Study Guide Unit 3 Periodicity Answer

1. Understanding the eBook Study Guide Unit 3 Periodicity Answer
 - The Rise of Digital Reading Study Guide Unit 3 Periodicity Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Unit 3 Periodicity Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Unit 3 Periodicity Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Unit 3 Periodicity Answer
 - Personalized Recommendations
 - Study Guide Unit 3 Periodicity Answer User Reviews and Ratings
 - Study Guide Unit 3 Periodicity Answer and Bestseller Lists
5. Accessing Study Guide Unit 3 Periodicity Answer Free and Paid eBooks

- Study Guide Unit 3 Periodicity Answer Public Domain eBooks
- Study Guide Unit 3 Periodicity Answer eBook Subscription Services
- Study Guide Unit 3 Periodicity Answer Budget-Friendly Options
- 6. Navigating Study Guide Unit 3 Periodicity Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Study Guide Unit 3 Periodicity Answer Compatibility with Devices
 - Study Guide Unit 3 Periodicity Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Unit 3 Periodicity Answer
 - Highlighting and Note-Taking Study Guide Unit 3 Periodicity Answer
 - Interactive Elements Study Guide Unit 3 Periodicity Answer
- 8. Staying Engaged with Study Guide Unit 3 Periodicity Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Unit 3 Periodicity Answer
- 9. Balancing eBooks and Physical Books Study Guide Unit 3 Periodicity Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Unit 3 Periodicity Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Unit 3 Periodicity Answer
 - Setting Reading Goals Study Guide Unit 3 Periodicity Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Unit 3 Periodicity Answer
 - Fact-Checking eBook Content of Study Guide Unit 3 Periodicity Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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