

Stoichiometry

Answer Key

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, dimensional analysis, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

[Note: The purpose of these questions is to activate prior knowledge and get students thinking. Students are not expected to know the answers to the Prior Knowledge Questions.]

1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? *176 grams of sugar*
2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? *15 sticks of butter*

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The Stoichiometry Gizmo™ allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

Fe_2O_3 *1*

CO *3*

Fe *2*

CO_2 *3*

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

Reactants

Iron: *2*

Oxygen: *6*

Carbon: *3*

Products

Iron: *2*

Oxygen: *6*

Carbon: *3*

View Equation

$$\text{Fe}_2\text{O}_3 (\text{s}) + 3\text{CO} (\text{g}) \rightarrow 2\text{Fe} (\text{s}) + 3\text{CO}_2 (\text{g})$$

View Solutions

How many moles of carbon monoxide (CO) are required to react completely with 1.75 moles of iron (III)-oxide (Fe_2O_3)?

Problems solved: 0

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Robert R. Redfield



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Teaching Science for Understanding James Joseph Gallagher, 2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities The Software Encyclopedia , 1988 **Subject Guide to Books in Print** , 1991

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workbook guides students through basic skills mathematical review and successful problem solving techniques Practice tests and solutions to odd numbered text problems are included **Study Guide with Selected Solutions** Karen

Timberlake,2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems **Principles of Modern Chemistry** Wade A. Freeman,David W. Oxtoby,1996 PRINCIPLES OF MODERN CHEMISTRY has dominated the honors and high mainstream general chemistry courses and is considered the standard for the course The fifth edition is a substantial revision that maintains the rigor of previous editions but reflects the exciting modern developments taking place in chemistry today Authors David W Oxtoby and H P Gillis provide a unique approach to learning chemical principles that emphasizes the total scientific process from observation to application placing general chemistry into a complete perspective for serious minded science and engineering students Chemical principles are illustrated by the use of modern materials comparable to equipment found in the scientific industry Students are therefore exposed to chemistry and its applications beyond the classroom This text is perfect for those instructors who are looking for a more advanced general chemistry textbook **Study Guide and Selected Solutions Manual for Fundamentals of**

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Corwin,Pearson,2004-07 Study Guide for Chemistry Jill Kirsten Robinson,John E. McMurry,Robert C. Fay,2019-01-04

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