

## BONDING REVIEW

1. An atom has 6 total electrons. How many electrons will be involved in bonding with another atom?
2. Intramolecular bonds are bonds within molecules while Intermolecular bonds are bonds between molecules.
3. What is Electronegativity?

④ only valence e<sup>-</sup> are involved in bonds

4. Fill in this table for Intramolecular bonds

TYPE OF BOND	Eneg Diff	Type of charges produced	2 EXAMPLES
Ionic	1.7 or greater	Strong	NaCl - any pair of atoms with
Polar covalent	0.5 - 1.7	slight $\delta^+$ $\delta^-$	KCl H <sub>2</sub> O, NH <sub>3</sub>
Non-polar covalent	less than 0.5	none	O <sub>2</sub> , Cl <sub>2</sub> , etc or any

5. For a molecule to be considered polar it needs to have:

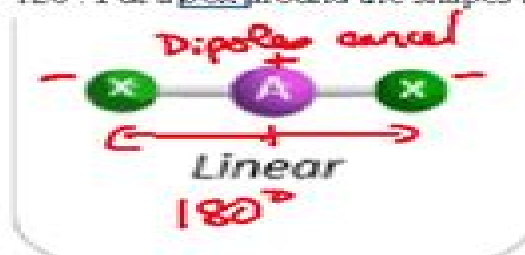
- i. polar bonds (Ionic or polar covalent)

and

- ii. Dipoles that DON'T cancel out

6. Look at the Molecular geometries below.

Circle the VSEPR shapes that could cause the dipoles to cancel out. Underline the shapes with a bond angle of 120°. Put a box around the shapes that would have lone pairs.



# Topic 6 Bonding Review Question Answers

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Rosenbrock**



## **Topic 6 Bonding Review Question Answers:**

## Whispering the Strategies of Language: An Mental Quest through **Topic 6 Bonding Review Question Answers**

In a digitally-driven world where screens reign great and immediate transmission drowns out the subtleties of language, the profound secrets and mental subtleties concealed within phrases usually go unheard. However, situated within the pages of **Topic 6 Bonding Review Question Answers** a fascinating fictional value pulsating with natural feelings, lies an extraordinary journey waiting to be undertaken. Written by a skilled wordsmith, that charming opus invites visitors on an introspective journey, lightly unraveling the veiled truths and profound influence resonating within the fabric of each word. Within the mental depths of this emotional evaluation, we will embark upon a heartfelt exploration of the book is primary styles, dissect its charming publishing model, and succumb to the strong resonance it evokes heavy within the recesses of readers hearts.

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