Name Key Period

Worksheet: Calculating Empirical & Molecular Formulas

1. The empirical formula for the compound having the formula $H_2C_2O_4$ is:

[A] C_2H_2 [B] CO_2H [C] COH [D] $C_2O_4H_2$ [E] COH_2

2. Calculate the empirical formula of a compound that is 85.6% C and 14.4% H (by mass).

[A] CH_2 [B] CH [C] C_2H_3 [D] C_2H_4 [E] C_2H

$C\ \frac{85.6}{12.01} = 7.13 \div 7.12 = 1$ $H\ \frac{14.4}{1.01} = 14.25 \div 7.12 = 2$

3. A compound is analyzed and found to contain 12.1% carbon, 16.2% oxygen, and 71.7% chlorine (by mass). Calculate the empirical formula of this compound.

[A] $COCl_2$ [B] CO_2Cl [C] CO_2Cl [D] $COCH$ [E] $COCl$

$C\ \frac{12.1}{12.01} = 1$ $Cl\ \frac{71.7}{35.45} = 2$

$O\ \frac{16.2}{16.00} = 1$

4. A compound contains 40.0% carbon, 6.7% hydrogen, and 53.3% oxygen (by mass). Calculate the empirical formula.

[A] C_2H_2O [B] CH_2O [C] CH_4O [D] C_2HO_2 [E] $C_2H_4O_2$

$C: \frac{40.0}{12.01} = 3.33 \div 3.33 = 1$ $O: \frac{53.3}{16.0} = 3.33 \div 3.33 = 1$

$H: \frac{6.7}{1.01} = 6.63 \div 3.33 = 2$

5. A compound contains 25.94% N and 74.06% O (by mass). What is the empirical formula?

$\frac{25.94\text{ g N}}{14.01} = \frac{1.84}{1.84} = 1 \times 2 = 2$

$\frac{74.06\text{ g O}}{16.00} = \frac{4.63}{1.84} = 2.5 \times 2 = 5$

N_2O_5

6. Determine the empirical formula of a compound containing 54.2% F and 45.8% S (by mass).

$\frac{54.2\text{ g F}}{19.00\text{ g/mol}} = \frac{2.85\text{ mol F}}{1.43} = 1.99 \sim 2$

$\frac{45.8\text{ g S}}{32.07\text{ g/mol}} = \frac{1.43\text{ mol S}}{1.43} = 1$

SF_2

7. A compound has 40.68% carbon, 5.12% hydrogen, and 54.20% oxygen (by mass). Calculate its empirical formula.

$\frac{40.68\text{ g C}}{12.01\text{ g/mol}} = \frac{3.387}{3.387} = 1 \times 2 = 2$ $\frac{5.12\text{ g H}}{1.01\text{ g/mol}} = \frac{5.069}{3.387} = 1.496 \times 2 = 3$ $\frac{54.20\text{ g O}}{16.00\text{ g/mol}} = \frac{3.387}{3.387} = 1 \times 2 = 2$

$C_2H_3O_2$

8. A 7.33-g sample of lanthanum, La, combines with oxygen to give 10.29 g of the oxide. Calculate the empirical formula of this oxide.

$10.29\text{ g} - 7.33\text{ g La} = 2.96\text{ g O}$

$\frac{7.33\text{ g La}}{139.91\text{ g/mol}} = \frac{.0527}{.0527} = 1 \times 2 = 2$ $\frac{2.96\text{ g O}}{16.00\text{ g/mol}} = \frac{.1850}{.0527} = 3.5 \times 2 = 7$

La_2O_7

9. Calculate the molecular formula of a compound with the empirical formula CH_2O and a molar mass of 150 g/mol.

30 $150 \div 30 = 5$

[A] $C_2H_4O_2$ [B] $C_5H_{10}O_5$ [C] $C_3H_6O_3$ [D] $C_6H_6O_6$ [E] $C_5H_{12}O_5$

Skill Practice 3moles And Formulas Answer Key

Lei Shi

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Table of Contents Skill Practice 3moles And Formulas Answer Key

1. Understanding the eBook Skill Practice 3moles And Formulas Answer Key
    - The Rise of Digital Reading Skill Practice 3moles And Formulas Answer Key
    - Advantages of eBooks Over Traditional Books
2. Identifying Skill Practice 3moles And Formulas Answer Key
    - Exploring Different Genres
    - Considering Fiction vs. Non-Fiction
    - Determining Your Reading Goals
3. Choosing the Right eBook Platform
    - Popular eBook Platforms
    - Features to Look for in an Skill Practice 3moles And Formulas Answer Key
    - User-Friendly Interface
4. Exploring eBook Recommendations from Skill Practice 3moles And Formulas Answer Key
    - Personalized Recommendations
    - Skill Practice 3moles And Formulas Answer Key User Reviews and Ratings
    - Skill Practice 3moles And Formulas Answer Key and Bestseller Lists
5. Accessing Skill Practice 3moles And Formulas Answer Key Free and Paid eBooks
    - Skill Practice 3moles And Formulas Answer Key Public Domain eBooks
    - Skill Practice 3moles And Formulas Answer Key eBook Subscription Services
    - Skill Practice 3moles And Formulas Answer Key Budget-Friendly Options
6. Navigating Skill Practice 3moles And Formulas Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- Skill Practice 3moles And Formulas Answer Key Compatibility with Devices
- Skill Practice 3moles And Formulas Answer Key Enhanced eBook Features

7. Enhancing Your Reading Experience
   - Adjustable Fonts and Text Sizes of Skill Practice 3moles And Formulas Answer Key
   - Highlighting and Note-Taking Skill Practice 3moles And Formulas Answer Key
   - Interactive Elements Skill Practice 3moles And Formulas Answer Key
8. Staying Engaged with Skill Practice 3moles And Formulas Answer Key
   - Joining Online Reading Communities
   - Participating in Virtual Book Clubs
   - Following Authors and Publishers Skill Practice 3moles And Formulas Answer Key
9. Balancing eBooks and Physical Books Skill Practice 3moles And Formulas Answer Key
   - Benefits of a Digital Library
   - Creating a Diverse Reading Collection Skill Practice 3moles And Formulas Answer Key
10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
11. Cultivating a Reading Routine Skill Practice 3moles And Formulas Answer Key
    - Setting Reading Goals Skill Practice 3moles And Formulas Answer Key
    - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Skill Practice 3moles And Formulas Answer Key
    - Fact-Checking eBook Content of Skill Practice 3moles And Formulas Answer Key
    - Distinguishing Credible Sources
13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
14. Embracing eBook Trends
    - Integration of Multimedia Elements
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