Chapter 2

Atom Components and Subatomic Particles

There is an attractive force between the negatively charged electrons and the positively charged nucleus.

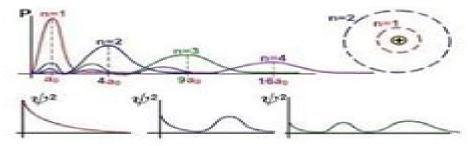
- The nucleus is smaller and contains most of the atom's mass.
- Relationship between the energy of an orbital and its average radius: As the radius increases, the orbital energy increases.
 - A small magnitude of E_n corresponds to a large denominator, which is n. So, if n is very large, then we are further from the nucleus: A weakly held electron.
 - As an electron gets farther from the nucleus, the principle quantum number (n) goes up and the energy increases and the electron is less stable.

Atomic Orbital Function Graphs

Ψ= wavefunction

 $|\Psi^2|$ = probability density

- Likelihood of locating the electron in a given region of space.
 - Form: y = sin(x)



Wavefunction describes orbitals

Determining an Orbital

N: Principle quantum number (shell)

- N=1, 2, 3...
- Determines the distance of an electron from the nucleus
- All orbitals with the same n (in the same atom) are the same distance from the nucleus on average.

L: Angular momentum quantum number (subshell)

- L= 0,1,2...(n-1)
- Determines the overall shape
- All orbitals in the same atom w/ the same n and I have the same angular momentum (describes shape)

Subatomic Physics Chapter Study Guide

Ernest M Henley, Stephen D Ellis

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