

Copper In Silver Nitrate Lab

by Charles Deaton

Background

The purpose of this lab is to find the theoretical percentage yield of the reaction. The reaction consisted of copper(Cu) and silver nitrate(AgNO_3), which after 3 class days of reacting, turned into copper nitrate($\text{Cu}(\text{NO}_3)_2$), and silver(Ag). In that third class day, they massed out that it had copper wire, and Ag and all silver nitrate solutions were measured. Once it was measured, the wire being over the side of a small beaker with a good portion of the wire submerged in AgNO_3 and all solutions. Observations were made and the three beakers were left out until day 3. When students returned a couple days later, and measured what they found to be a beaker. Stripping their beaker the wire out of the solution and submerging with water so that the precipitate were over the beaker. This mass of the wire was taken again, and a piece of filter paper was also weighed. The filter paper was then set up in a dry beaker with a funnel and a large beaker beside. The concentrated HNO_3 beaker was poured through the filter paper, and mixed with more water. This was to remove all of the reactants from the beaker. Day 4 was to remove the final mass of the products so that filter paper that could be dry from the days prior. After everything was measured and written in the table, calculations were done to find the theoretical yield, percent yield, and percent error of the lab.

Silver Nitrate Lab Report

Joacim Rocklöv

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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the fundamentals covered in other books focusing instead on practical applications of electroanalysis **Bibliography of Scientific and Industrial Reports** ,1965-07 **U.S. Government Research & Development Reports** ,1971

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