

1. (a) Le Chatelier's Principle:

If a system in equilibrium is subjected to a constraint, a process will occur (1) which tends to counteract the constraint. /AW

(b) (i) While lower temperatures than  $450^{\circ}\text{C}$  would give a high yield of  $\text{SO}_3$ , they would make the reaction rate too slow. /AW (1)

(ii) Cooling is necessary to favour the forward exothermic reaction. /AW (1)

(iii) Cost of thicker pipes to withstand the higher pressures will be too high. /AW (1)

(c)  $K_p = \frac{p}{p_{\text{SO}_2} p_{\text{O}_2}}$  (1)

units =  $\frac{\text{atm}}{(\text{atm})^2} = \frac{1}{\text{atm}}$  (1)

(d) (i) No change in  $K_p$  value. (1)

Increase in partial pressure of  $\text{SO}_2$  is counter-balanced by increase in partial pressure of  $\text{O}_2$ .

Extra  $\text{SO}_2$  reacts to produce more  $\text{SO}_3$  until equilibrium is re-established. /AW (1)

(ii) Addition of  $\text{V}_2\text{O}_5$  catalyst has no effect on  $K_p$ . (1)

Catalyst only increases rate of attaining equilibrium, without changing proportions of reactants and products. /AW (1)

(iii)  $K_p$  value decreases. (1)  
Endothermic decomposition of  $\text{SO}_3$  to form more  $\text{SO}_2$  and  $\text{O}_2$  is favoured, thus lowering  $K_p$ . (1)

[Total: 12]

A)  $K_p = \frac{p}{(p_{\text{SO}_2})^2 (p_{\text{O}_2})}$

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