

NAME: _____ EQUILIBRIUM PRACTICE

Equilibrium Practice

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1. The reaction below has reached a state of equilibrium. Write the equilibrium constant expression for this reaction.
2. Consider the following chemical equation: $2\text{H}_2\text{O}(g) \rightleftharpoons 2\text{H}_2(g) + \text{O}_2(g)$. In equilibrium, the concentration of H_2 is 0.02 M, the concentration of H_2O is 0.01 M, and the concentration of O_2 is 0.01 M. Calculate the equilibrium constant, K .
3. At a certain pressure, 10 g of each of the gases, chlorine and hydrogen, were placed in a 100 L container. Calculate the equilibrium constant, K , for the following equation: $\text{H}_2 + \text{Cl}_2 \rightleftharpoons 2\text{HCl}$. At equilibrium, the concentration of each gas was 0.01 M.
4. A 100 mL flask is filled with a gas mixture at a total pressure of 1 atm. The flask is sealed and the gas mixture is allowed to reach equilibrium. Calculate the equilibrium constant, K , for the reaction: $\text{H}_2 + \text{Cl}_2 \rightleftharpoons 2\text{HCl}$.
5. Consider the following chemical equation: $\text{PCl}_5(g) \rightleftharpoons \text{PCl}_3(g) + \text{Cl}_2(g)$. At the initial concentration of PCl_5 is 0.10 M, what is the equilibrium concentration of Cl_2 if the concentration of PCl_3 is 0.05 M?

Skill Practice 45 Equilibrium Practice Answers

Rod Powers



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