Name:

Use the periodic table to find the two ions (positive ion and negative ion) for each compound. Then write the formula.

Remember, the overall charge of the compound is neutral (zero charge). Thus, the overall positive charge must be equal to the overall negative charge.

Example: calcium chloride

 Ca^{2+} Cl^- (These two charges are not the same, but 2 chloride ions will give an overall negative charge of -2. The positive ion already has a + 2 charge.)

Ca2+Cl2

Ca²*

Count up the total charge. +2, -1, -1 This adds up to zero!!

Therefore, CaCl2 is the answer.

Name	Positive Ion	Negative Ion	Formula
1) sodium chloride ex.	Na *	ci-	NaCl
2) potassium bromide	-		
3) calcium oxide			
4) magnesium sulfide			
5) zinc chloride			
6) sodium sulfide	-		
7) strontium fluoride	1.7		
8) lithium oxide			
9) potassium nitride			
10) sodium phosphide			
II) aluminum chloride			
(2) calcium nitride			
13) aluminum oxide		T T	
(4) magnesium phosphide			
l5) aluminum sulfide			

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