

9.1

THE ARITHMETIC OF EQUATIONS

SECTION REVIEW

Objectives

- Calculate the amount of reactants required or product formed in a chemical process.
- Interpret balanced chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

Key Terms

- stoichiometry

Part A Completion

Copy this completion exercise for class. Your understanding of the concepts and terms is clear when you can do this exercise. Blank blanks cover the completed work at home, school, library, or home first.

- The coefficients in a balanced chemical equation indicate _____ 1. _____ the relative number of _____ 2. _____ of reactants and products. All _____ 3. _____ stoichiometric calculations begin with a _____ 4. _____ On a _____ 5. _____ and _____ 6. _____ are conserved in every reaction: mass, volume, and representative particles may not be.
- In solving stoichiometric problems, conversion factors relating moles of reactants to _____ 7. _____ of products are used.
- If you assume _____ 8. _____, the equation also tells you about the volume of gas.

Part B True-False

(Quickly mark off the statements as to be always true, AC; sometimes true, ST; or never true, NT.)

- _____ 7. The coefficients in a balanced chemical equation can be used to determine the relative limiting reactants to products.
- _____ 8. The coefficients in a balanced chemical equation tell the relative volumes of reactants and products, expressed in any suitable units of volume.
- _____ 9. To calculate the mass of a molecule in grams, you can use the molar mass and Avogadro's number.

The Arithmetic Of Equations Section 9 Answer Key

Elise Baniam, Michael Smith



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