

WRITING CHEMICAL EQUATIONS

Name _____	Section _____	Date _____
Writing Equations		
Directions: Write and balance each chemical equation. You will need to write the products. Include phase labels with each reactant and product. You may need to refer to a table of solubility rules.		
1) Aqueous sulfuric acid reacts with aqueous potassium hydroxide		
2) Aqueous potassium carbonate reacts with aqueous hydrochloric acid		
3) Hydrogen and oxygen gases combine to form water vapor		

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1) Aqueous sulfuric acid reacts with aqueous potassium hydroxide		
2) Aqueous potassium carbonate reacts with aqueous hydrochloric acid		
3) Hydrogen and oxygen gases combine to form water vapor		
4) Methane gas (carbon tetrahydride) reacts with chlorine gas to yield chloroform (CH ₃ Cl) gas and hydrochloric acid gas		
5) Aqueous ammonium sulfide reacts with aqueous magnesium sulfate to yield solid magnesium sulfide and aqueous ammonium sulfate		
6) Mercury (II) nitrate reacts with aqueous iron(III) iodide to form aqueous iron (III) nitrate and solid mercury (II) iodide		
7) Aqueous strontium hydroxide reacts with aqueous iron (III) nitrate to produce solid iron hydroxide and aqueous strontium nitrate		
8) Solid zinc carbonate reacts with aqueous potassium chloride. An aqueous solution of zinc chloride and potassium carbonate is formed.		
9) Aqueous ammonium chromate reacts with aqueous barium sulfide react to form aqueous ammonium sulfide and solid barium chromate.		
10) Aqueous lithium oxalate reacts with aqueous barium sulfate to form aqueous lithium sulfate and solid barium oxalate		

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A CHEMISTRY WORKSHEET

10 PROBLEMS

WORKED ANSWER KEY INCLUDED

Topic 3 Formulas And Equations Answer Key

Lyn Baker



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