



# Oxidation-Reduction Reactions

- Electrons can be neither created out of nothing nor destroyed
- In any redox reaction there is an element being reduced and an element being oxidized
- The total increase in the oxidation numbers must equal the total decrease in the oxidation numbers

If an element is reduced

- It gains electrons
- Oxidation number decreases
- The substance is classified as an oxidizing agent

If an element is oxidized

- It loses electrons
- Oxidation number increases
- The substance is classified as a reducing agent

# Teaching Transparency 56 Oxidation And Reduction Reactions

**Ruren Xu,Yan Xu**



## Teaching Transparency 56 Oxidation And Reduction Reactions:

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**Demonstrations in Chemistry** George L. Gilbert, 1994

**The Popular Science News and Boston Journal of**

**Chemistry** ,1869 *Boston Journal of Chemistry* ,1871 *Chemical Age* ,1928 *The Chemical Age* ,1928 [Index to](#)

[35mm Educational Filmstrips](#) National Information Center for Educational Media, 1975

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[Modern Inorganic Synthetic Chemistry](#) Ruren Xu, Yan Xu, 2017-02-11 Modern

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