

43. Electropositivity of elements across the periodic table normally

A. remains constant down the group.
 B. increases across the period.
 C. decreases across the period.
 D. decreases down the group.

44. Consider the following table:

Substance	Melting point / °C	Boiling point / °C
P	-78	-25
Q	-8	40
R	-6	150
S	44	289

Which of the substances is a liquid at room temperature and rapidly evaporates on exposure to air?

A. P
 B. Q
 C. R
 D. S

45. Which of the following oxides is amphoteric?

A. Carbon(II) oxide
 B. Nitrogen(IV) oxide
 C. Lead(II) oxide
 D. Calcium oxide

46. Which of the following processes occur during fractional distillation of petroleum?

A. Condensation and diffusion
 B. Diffusion and evaporation
 C. Diffusion and sublimation
 D. Evaporation and condensation

47. If the molar mass of $X(\text{HCO}_3)_2$ is 162 g mol^{-1} , determine the relative atomic mass of X. [H = 1.0, C = 12.0, O = 16.0]

A. 40
 B. 48
 C. 61
 D. 101

48. Charcoal is used in the decolourization of sugar because of its

A. absorption property.
 B. amorphous property.
 C. oxidizing property.
 D. adsorption property.

49. The first definition of an element was made by

A. J. Dalton.
 B. A. Lavoisier.
 C. R. Boyle.
 D. J. J. Thompson.

50. Which of the following scientists formulated the law of conservation of mass?

A. A. Lavoisier
 B. J. Dalton
 C. R. Boyle
 D. J. Proust

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Science Chemistry ,2024 Chemistry Phillip John Buddle,J. S. Mackay,1970

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