

- 4 a Calcium ions are 2+ and chloride ions are 1-, so the formula is CaCl₂.
 - b Silver ions are 1+ and nitrate ions are 1-, so the formula is AgNO₃.
 - c Copper(s) is Cu²⁺ and phosphate is PO₄³⁻. For a neutral compound you need three copper ions and two phosphate ions, making charges of +6 and -6. Therefore, the formula is Qu₂(PO₄)₃₋.
 - d Aluminium is 3+ and oxide is 2-, so aluminium oxide is Al₂O₃.
 - There are two ways of working out formulae of ionic compounds. One way is to make sure that the total positive charge equals the total negative charge (see part c). The other way, if the charges are not the same, is to swap them round. Copper is 2+ and phosphate 3-, so there are three copper ions and two phosphate ions in the formula.
- 5 a As there is one manganese ion to each carbonate ion, the charges have the same numerical value. Carbonate ions are 2-, so the manganese ions in this compound must be 2+.
 - b The charge on two vanadium ions must equal the charge on three sulfate ions. Each sulfate is -2, so three sulfate ions have a total charge of -6. This means that two vanadium ions have a total charge of +6, so each has a charge of +3.
- 6 a molar mass of Ca(OH)₂ = 40.1 + 2 × (16.0 + 1.0) = 74.1 g mol⁻¹
 - b molar mass of $Al_2(SO_4)_1 = (2 \times 27.0) + 3 \times [32.1 + (4 \times 16.0)] = 342.3 \text{ g mol}^{-2}$
 - c molar mass of FeSO₄.7H₂O = $55.8 + 32.1 + (4 \times 16.0) + 7 \times (2.0 + 16.0) = 277.9 \text{ gmol}^{-1}$
 - ② In calculating molar mass, make sure that you use the relative atomic masses (the larger numbers), not the atomic numbers, from the periodic table. Molar mass has the units of g mol⁻¹. Relative molecular mass has no units. The numbers are the same.
- 7 The balanced equations are:
 - a $P_4 + 5O_2 \rightarrow P_4O_{23}$ or $2P_2O_5$
 - Phosphorus(v) oxide exists as the dimer P₂O₁₀, but 2P₂O₃ on the right of the equation would be acceptable.
 - b 4AI + 3O₂ → 2AI₂O₂
 - c $2Mg(NO_3)_2 \rightarrow 2MgO + 4NO_2 + O_3$
 - In part c, balance the oxygen atoms last. Remember that the number of oxygen atoms on the right-hand side of the equation must be an even number, because Mg(NO₂)₂ contains an even number of oxygen atoms.
 - d 2LiOH + H₂SO₄ → Li₂SO₄ + 2H₂O
 - $e^{-}2Fe^{3+} + 2f^{-} \rightarrow 2Fe^{2+} + I_{-}$
 - Make sure that this equation balances for charge. Having one Fe³⁺ ion on the left-hand side and one Fe³⁺ on the right-hand side balances the number of Felions, but the charge on the left would not equal the charge on the right.
 - $f = 2C_0H_{18} + 25O_2 \rightarrow 16CO_2 + 18H_2O$ or $C_0H_{18} + 12\frac{1}{2}O_2 \rightarrow 8CO_2 + 9H_2O$
 - The only source of oxygen on the left-hand side of the equation is O₂, so the number of oxygen atoms on the right-hand side must be an even number. Therefore, theremust be an even number of H₂O molecules on the right. This requires two C₂H₁₂ molecules on the left.
- 8 For equations a-c, the simplest way is to realise that the component ions of the compound combine to give the precipitate. All other ions are spectators.

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John S. Phillips, Cheryl Wistrom

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