

CP6 Radioactivity

Background Radiation

What is natural background radiation?

State 2 sources of natural background radiation:

Draw a labelled diagram of Rutherford's model State 2 sources of background radiation that does not occur naturally

How can radiation be measured and describe how the equipment works.

Describe the experiment Rutherford carried out which allowed him to come up with his model.

Types of radiation

Describe the structure of an alpha particle What is the radius of a nucleus?

What is the radius of an atom?

What is the charge on an alpha particle?

How is Rutherford's model of the atom similar to Bohr's model? Describe the structure of a beta particle

What is the charge on a beta particle?

Inside Atoms

Draw a labelled diagram of the modern day structure of an atom Describe the structure of a gamma ray

What does unstable mean?

	Mass	Charge
Proton		
Neutron		
Electron	Alpha	Beta
Gamma		Weak

Complete the table

Penetration	C	N	Mg	Co
Mass number				
Atomic number				
Protons				
Neutrons				

Explain why alpha particles are more ionising than beta particles

Define an isotope

Radioactive decay

Complete the table

Particle	Symbol
Alpha	
Beta	β^-
Positron	
Neutron	n

Describe how a magnesium atom becomes a Mg^{2+} ion. What happens to the atomic and mass number when a neutron is emitted?

Electrons and Orbits

Radioactive decay happens when an atom absorbs electromagnetic radiation Describe what happens during β^- decay and the effect on the nucleus

Describe what happens when an atom emits -- electromagnetic radiation Describe what happens during β^- decay and the effect on the nucleus

What is the difference between emission and absorption spectra Complete the table summarising what happens to the mass and atomic number during radioactive decay

	Mass number	Atomic number
Alpha	Decreases by 4	Decreases by 2
Beta	Increases by 0	Increases by 1

Radium has a mass number of 226 and an atomic number of 88.

Write a nuclear equation when this nucleus emits an alpha particle

Write a nuclear equation if Iodine-131 undergoes β^- decay. Iodine has an atomic number of 53 and loses an electron to form an ion. How many protons and electrons are there? Write a nuclear equation if Potassium-37 undergoes β^- decay.

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Didier Musso



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