

Modeled after AP 2013 #21:

- $$2 \text{C}_2\text{H}_6\text{g} + 7 \text{O}_2\text{g} \rightarrow 4 \text{CO}_2\text{g} + 6 \text{H}_2\text{Og}$$

In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of CO_2g is produced?

 - 0.50 mol
 - 1.50 mol
 - 4.50 mol**
 - 5.00 mol
- $$\text{C}_2\text{H}_6\text{g} + 5 \text{O}_2\text{g} \rightarrow 4 \text{CO}_2\text{g} + 6 \text{H}_2\text{Og}$$

In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of CO_2g is produced?

 - 0.50 mol
 - 1.00 mol
 - 3.50 mol**
 - 7.00 mol
- $$2 \text{C}_2\text{H}_6\text{g} + 7 \text{O}_2\text{g} \rightarrow 4 \text{CO}_2\text{g} + 6 \text{H}_2\text{Og}$$

In the reaction represented above, what is the total number of moles of products made when 3.50 mole of O_2g is consumed?

 - 2.00 mol
 - 3.00 mol
 - 5.00 mol**
 - 10.0 mol
- $$\text{C}_2\text{H}_6\text{g} + 5 \text{O}_2\text{g} \rightarrow 4 \text{CO}_2\text{g} + 6 \text{H}_2\text{Og}$$

In the reaction represented above, what is the total number of moles of products made when 3.50 mole of O_2g is consumed?

 - 1.00 mol
 - 2.00 mol**
 - 4.00 mol
 - 8.00 mol

Modeled after AP 2014g #48:

- $$4 \text{NH}_3\text{g} + 5 \text{O}_2\text{g} \rightarrow 4 \text{NOg} + 6 \text{H}_2\text{Og}$$

A 7.0 mol sample of NH_3g and a 10. mol sample of O_2g are placed in a previously evacuated container and react according to the equation above. How many moles of NOg are produced when one of the reactants has been totally consumed?

 - 2.5 mol
 - 7.0 mol**
 - 8.0 mol
 - 28 mol
- $$\text{C}_2\text{H}_6\text{g} + 3 \text{H}_2\text{g} \rightarrow 2 \text{CH}_4\text{g}$$

Samples containing 4.0 mol $\text{C}_2\text{H}_6\text{g}$ and 5.0 mol H_2g are placed in a previously evacuated container and allowed to react according to the equation above. When one of the reactants is completely consumed, how much product is formed?

Zumdahl Ap Chemistry Review Questions Answers

Mina Lebitz



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