

AP Chemistry Chapter 4 Answers – Zumdahl

4.36

The solubility rules referenced in the following answers are from Table 4.1 of the text. The phrase “slightly soluble” is interpreted to mean insoluble and the phrase “marginally soluble” is interpreted to mean soluble.

- a. Soluble (Rule 3)
- b. Soluble (Rule 1)
- c. Insoluble (Rule 4)
- d. Soluble (Rules 2 and 3)
- e. Insoluble (Rule 6)
- f. Insoluble (Rule 5)
- g. Insoluble (Rule 6)
- h. Soluble (Rule 2)

4.38

Use Table 4.1 to predict the solubility of the possible products.

- a. Possible products = Hg_2SO_4 and $\text{Cu}(\text{NO}_3)_2$; precipitate = Hg_2SO_4
- b. Possible products = NiCl_2 and $\text{Ca}(\text{NO}_3)_2$; Both salts are soluble so no precipitate forms.
- c. Possible products = KI and MgCO_3 ; precipitate = MgCO_3
- d. Possible products = NaBr and $\text{Al}_2(\text{CrO}_4)_3$; precipitate = $\text{Al}_2(\text{CrO}_4)_3$

4.44

- a. $\text{CrCl}_3(\text{aq}) + 3\text{NaOH}(\text{aq}) \rightarrow \text{Cr}(\text{OH})_3(\text{s}) + 3\text{NaCl}(\text{aq})$
 $\text{Cr}^{3+}(\text{aq}) + 3\text{OH}^-(\text{aq}) \rightarrow \text{Cr}(\text{OH})_3(\text{s})$
- b. $2\text{AgNO}_3(\text{aq}) + (\text{NH}_4)_2\text{CO}_3(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s}) + 2\text{NH}_4\text{NO}_3(\text{aq})$
 $2\text{Ag}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s})$
- c. $\text{CuSO}_4(\text{aq}) + \text{Hg}_2(\text{NO}_3)_2(\text{aq}) \rightarrow \text{Cu}(\text{NO}_3)_2(\text{aq}) + \text{Hg}_2\text{SO}_4(\text{s})$
 $\text{Hg}_2^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) \rightarrow \text{Hg}_2\text{SO}_4(\text{s})$
- d. No reaction occurs because all possible products (SrI_2 and KNO_3) are soluble.

4.46

Because no precipitates formed upon addition of NaCl or Na_2SO_4 , we can conclude that Hg_2^{2+} and Ba^{2+} are not present in the sample since Hg_2Cl_2 and BaSO_4 are insoluble salts. However, Mn^{2+} may be present since Mn^{2+} does not form a precipitate with either NaCl or Na_2SO_4 . A precipitate formed with NaOH ; the solution must contain Mn^{2+} because it forms a precipitate with OH^- [$\text{Mn}(\text{OH})_2(\text{s})$].

4.47



$$0.0750 \text{ L} \times 0.100 \text{ mol AgNO}_3/1 \text{ L} \times 1 \text{ mol Na}_2\text{CrO}_4/2 \text{ mol AgNO}_3 \times$$

$$161.98 \text{ g Na}_2\text{CrO}_4/1 \text{ mol Na}_2\text{CrO}_4 = 0.607 \text{ g Na}_2\text{CrO}_4$$

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