AP Chemistry Chapter 4 Answers - Zumdahl

4.36

The solubility rules referenced in the following answers are from Table 4.1 of the text. The phrase "slightly soluble" is interpreted to mean insoluble and the phrase "marginally soluble" is interpreted to mean soluble.

- Soluble (Rule 3)
- b. Soluble (Rule 1)
- c. Insoluble (Rule 4)
- Soluble (Rules 2 and 3)
- e. Insoluble (Rule 6)
- Insoluble (Rule 5)
- g. Insoluble (Rule 6)
- h. Soluble (Rule 2)

4.38

Use Table 4.1 to predict the solubility of the possible products.

- Possible products = Hg₂SO₄ and Cu(NO₁)₅; precipitate = Hg₂SO₄
- Possible products = NiCl₂ and Ca(NO₃)₂; Both salts are soluble so no precipitate forms.
- Possible products = KI and MgCO₃; precipitate = MgCO₃
- d. Possible products = NaBr and Al₂(CrO₄)₃; precipitate = Al₂(CrO₄)₃

4.44

a. $CrCl_3(aq) + 3NaOH(aq) \rightarrow Cr(OH)_3(s) + 3NaCl(aq)$

$$Cr^{3*}(aq) + 3OH'(aq) \rightarrow Cr(OH)_3(s)$$

b. $2AgNO_3(aq) + (NH_4)_2CO_3(aq) \rightarrow Ag_2CO_3(s) + 2NH_4NO_3(aq)$

$$2Ag^*(aq) + CO_3^2(aq) \rightarrow Ag_2CO_3(s)$$

c. $CuSO_4(aq) + Hg_2(NO_3)_2(aq) \rightarrow Cu(NO_3)_2(aq) + Hg_2SO_4(s)$

$$Hg_2^{2*}(aq) +_-SO_4^{2*}(aq) \rightarrow Hg_2SO_4(s)$$

No reaction occurs because all possible products (SrI₂ and KNO₃) are soluble.

4.46

Because no precipitates formed upon addition of NaCl or Na₂SO₄, we can conclude that Hg₂^{2*} and Ba^{2*} are not present in the sample since Hg₂Cl₂ and BaSO₄ are insoluble salts. However, Mn^{2*} may be present since Mn^{2*} does not form a precipitate with either NaCl or Na₂SO₄. A precipitate formed with NaOH; the solution must contain Mn^{2*} because it forms a precipitate with OH [Mn(OH)₂(s)].

4.47

 $2AgNO_3(aq) + Na_2CrO_4(aq) \rightarrow Ag_2CrO_4(s) + 2NaNO_3(aq)$

0.0750 L x 0.100 mol AgNO₃/1 L x 1 mol Na₂CrO₄/2 mol AgNO₃ x

161.98 g Na₂CrO₄/1 mol Na₂CrO₄ = 0.607 g Na₂CrO₄

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