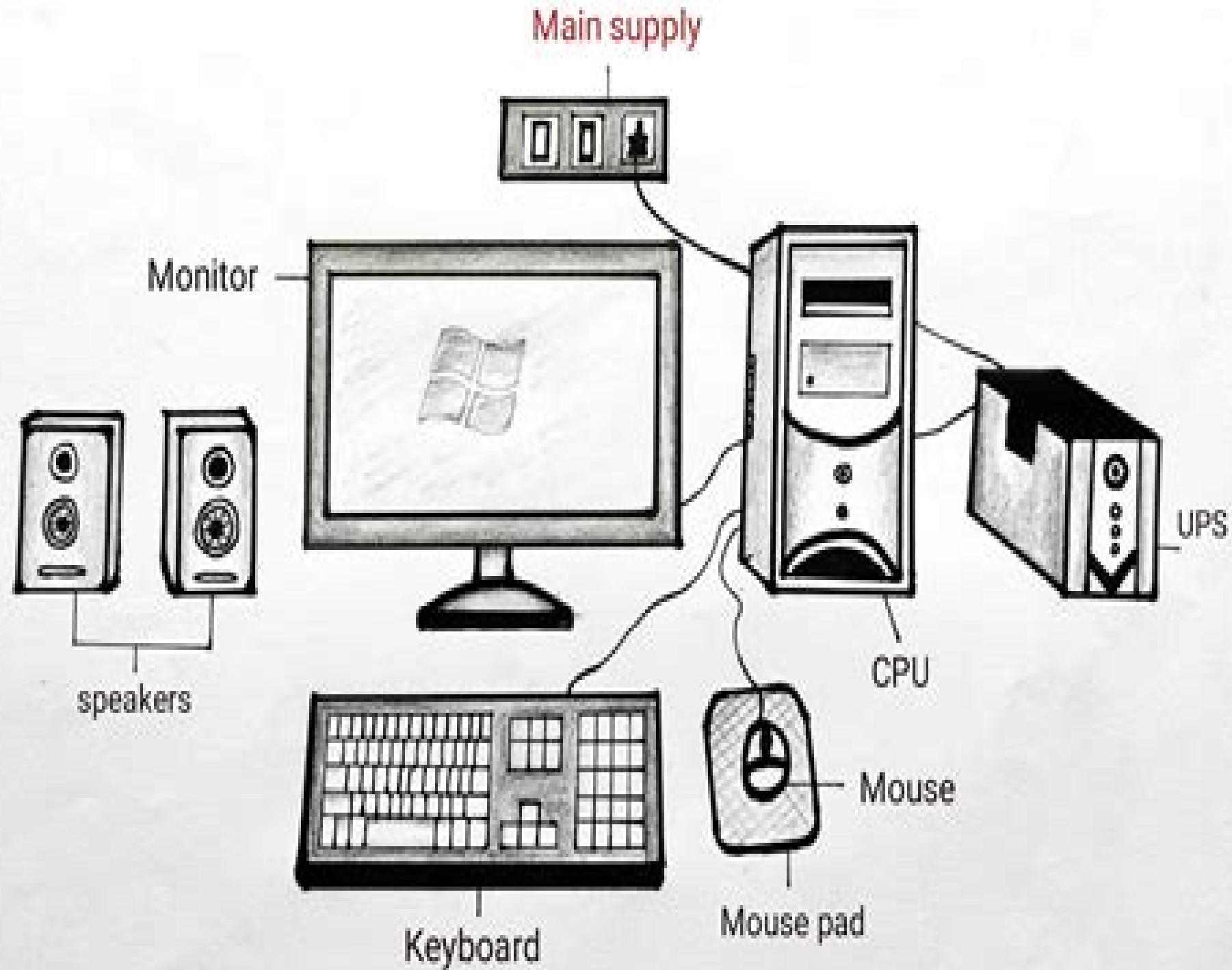




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thermodynamic equilibrium 4L C System of variable composition Partial molar quantities chemical potential Van t Hoff s reaction isotherm relation between standard free energy change equilibrium constant thermodynamic derivation of law of mass action effect of temperature on equilibrium constant reaction isochore 4L Unit III Phase Equilibria A Phase rule Statement of phase rule definition of phase component and degree of freedom derivation of phase rule Clapeyron equation its application in deciding slopes of line for two phase equilibria applications of phase rule to two phase equilibria of i water system ii sulphur system iii Pb Ag system 6L B Liquid Liquid mixtures Ideal liquid mixtures Raoult's law of ideal solutions Henry's law non ideal systems azeotropes HCl H₂O ethanol water system Partial miscible liquids Phenol water system trimethylamine water nicotine water system lower upper consolute temperature effect of impurity Immiscible liquids Steam distillation Nernst distribution law Limitations deviations applications 6L Unit IV Solid State Laws of crystallography i Law of constancy of interfacial angles ii Law of rationality of indices iii Law of symmetry symmetry of elements in crystals Unit cell space lattice orientation of lattice plane Miller indices Bravais lattices crystal systems X ray diffraction by crystal derivation of Bragg's equation

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